

Annotated periodic table of the elements

1	H 1.01 Hydrogen		Group name*																He 4.00 Helium			
	2	1 Hydrogen & the alkali metals 2 Alkaline earth metals 11 Coinage metals 12 Volatile metals 13 Lanthanides																14 Crystallogens 15 Pnictogens 16 Chalcogens 17 Halogens 18 Helium group				
	*Groups 1 to 10 are named after their first members i.e. Group 1 is the Scandium group																					
1	3 Li 6.94 Lithium		4 Be 9.01 Beryllium		Most common oxidation states: electronegativity (Mulliken-Pauling)																10 Ne 20.18 Neon	
2	11 Na 22.99 Sodium		12 Mg 24.31 Magnesium																		18 Ar 39.95 Argon	
3	19 K 39.10 Potassium		20 Ca 40.08 Calcium																		36 Kr 83.80 Krypton	
4	37 Rb 85.47 Rubidium		38 Sr 87.62 Strontium																		84 Po 209 Polonium	
5	55 Cs 132.91 Cesium		56 Ba 137.33 Barium																		118 Og 284.15 Oganesson	
6	87 Fr 223.02 Francium		88 Ra 226.03 Radium																			
7	117 Ts 289.10 Tennessine		118 Og 284.15 Oganesson																			

- Group names*
- Hydrogen & the alkali metals
 - Alkaline earth metals
 - Coinage metals
 - Volatile metals
 - Isotopes
 - Chalcogens
 - Postcogens
 - Chalcogens
 - Halogens
 - Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

Main group
oxidation states
electronegativity
Mendeleev

Atomic number
and symbol
Atomic weight
and name

CLASS COUNTERPARTS

Metals	Nonmetals
Pre-transition	
Active	Active
Transition	Intermediate
Post-transition	Metals
Metals (Ru-Pt, Os-Au)	Metals

Active metals

- Pre-transition metals**
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).
- Rare earths**
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.
- Actinides**
Soft, dense and reactive metals. Those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals

- Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.
- Noble metals**
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very easily electroplated; insoluble oxides.
- Frontier metals**
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and bismuthates.

Metals

- Metalloids**
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750-1,000 kJ/mol), and electronegativity values (1.8-2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.
- Intermediate nonmetals**
H and N are colorless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).
- Group 1**
The alkali metals Li, Na, K, Rb, Cs, and Fr are soft, silvery-white, and highly reactive. They are found in nature as compounds, not as free elements. They are the most reactive of the elements.

Nonmetals

- Chalcogens**
O, S, Se, and Te are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.
- Noble gases**
Colorless, odorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Radioactive
Oganesson is likely to be either a metalloid or a post-transition metal

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L Reisser



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1. Understanding the eBook Tour Of The Periodic Table Answers
 - The Rise of Digital Reading Tour Of The Periodic Table Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Tour Of The Periodic Table Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Tour Of The Periodic Table Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Tour Of The Periodic Table Answers
 - Personalized Recommendations
 - Tour Of The Periodic Table Answers User Reviews and Ratings
 - Tour Of The Periodic Table Answers and Bestseller Lists
5. Accessing Tour Of The Periodic Table Answers Free and Paid eBooks
 - Tour Of The Periodic Table Answers Public Domain eBooks
 - Tour Of The Periodic Table Answers eBook Subscription Services
 - Tour Of The Periodic Table Answers Budget-Friendly Options
6. Navigating Tour Of The Periodic Table Answers eBook Formats

- ePub, PDF, MOBI, and More
- Tour Of The Periodic Table Answers Compatibility with Devices
- Tour Of The Periodic Table Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Tour Of The Periodic Table Answers
 - Highlighting and Note-Taking Tour Of The Periodic Table Answers
 - Interactive Elements Tour Of The Periodic Table Answers
- 8. Staying Engaged with Tour Of The Periodic Table Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Tour Of The Periodic Table Answers
- 9. Balancing eBooks and Physical Books Tour Of The Periodic Table Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Tour Of The Periodic Table Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Tour Of The Periodic Table Answers
 - Setting Reading Goals Tour Of The Periodic Table Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Tour Of The Periodic Table Answers
 - Fact-Checking eBook Content of Tour Of The Periodic Table Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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