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Chapter 5 The Periodic Rable

Investigation SA

Using Clues to Identify Elements

Background Information

Chemical elements can be classified according to their properties as metals, normetals, and metalloids. Metals are good conductors of heat and electricity. Many metals are mallinable and ductile. Nonmetals are poor conductors of heat and electricity, and solid normetals tend to be brittle. Metalloids have properties between those of metals and nonmetals.

Elements in the same group on the periodic table have the same number of valence electrons, which are electrons in the highest occupied energy level of an atom. The number of an A group matches the number of valence electrons in atoms of each element in the group. For example, the Group 4A elements each have four valence electrons per atom. The exception to this pattern is the element helium, which is in Group 8A leat has only two valence electrons.

Because elements in a group have the same reamber of valence electrons, they tend to have similar properties. The most reactive metals are the alkali metals in Group IA on the far left side of the periodic table. The Group IA elements are the alkaline earth metals, which are somewhat less reactive than the alkali metals. Groups of elements become less metallic in their properties from left to right across the periodic table. The most reactive more from left to right across the periodic table. The most reactive more from the truth are the colories and odorless, and narely react with other elements.

In this investigation, you will use a list of closes to identify 34 elements and place them in their correct locations in the periodic table.

Problem

Where do the elements described in the class fit in the periodic table?

Pre-Lab Discussion

Bend the entire investigation. Then, work with a partner to answer the following questions.

Analyzing Data. How will the index cards be useful when you are reading the clues to the elements?		
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