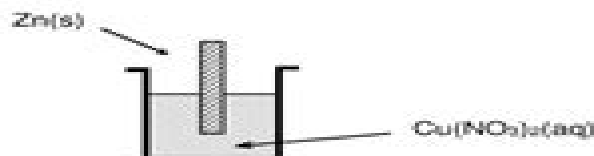


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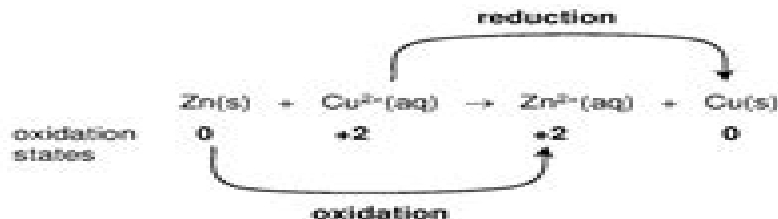
In this module, we turn our attention to **oxidation-reduction reactions**.

e.g.



◀ To identify an oxidation-reduction reaction, you must assign oxidation states. (Review rules for assigning oxidation states, if necessary.) Keep in mind that the oxidation state is the hypothetical charge an atom would have if the bonding electrons in each bond were transferred to the more electronegative atom.

A copper coating forms spontaneously on the zinc rod because of the following reaction.



Think OILRIG.

Oxidation
Is
Loss of e^-
Reduction
Is
Gain of e^-

All oxidation-reduction reactions can be "split" into two separate half-reactions, as shown below.



In this particular example, the electron transfer is "direct" in the sense that the reactants are in direct contact with each other. Note that the electrons don't appear in the overall chemical equation. The interesting (and useful) thing about electron transfer reactions is that they can be carried out even if the reactants are not in direct contact!

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