

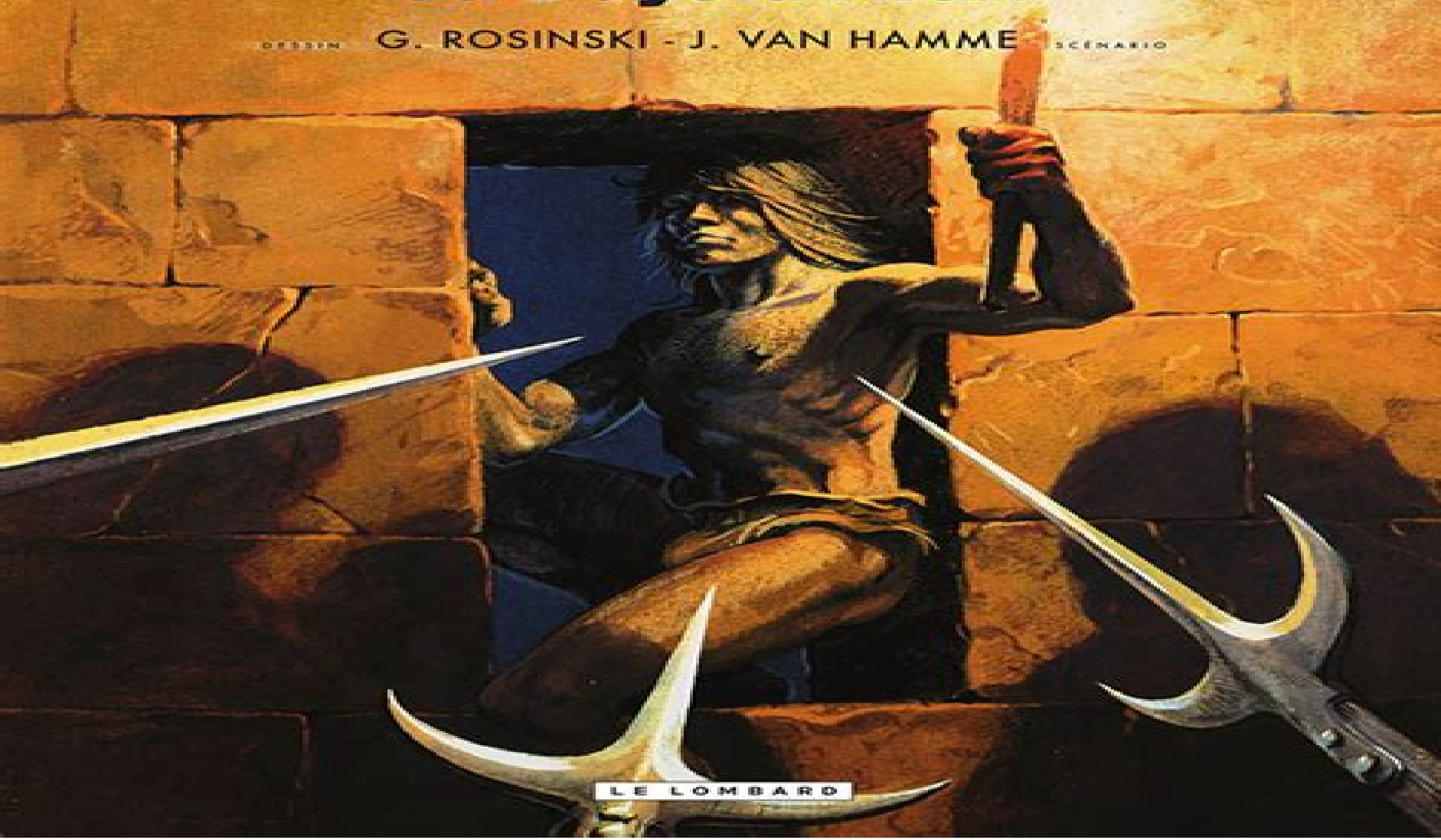


THORCAL



Les 3 Vieillards du Pays d'Aran

DESSIN G. ROSINSKI - J. VAN HAMME SCÉNARIO



LE LOMBARD

Thorgal Tome Oles Trois Vieillards Du Pays Daran

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+ \rightarrow 2\text{AgCl}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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$1\text{K}_3\text{PO}_4 + \text{Al}(\text{OH})_3 + 3\text{HNO}_3 \rightarrow 3\text{KNO}_3 + \text{AlPO}_4 + 3\text{H}_2\text{O}$ net ionic $\text{Al}^{3+} + 3\text{PO}_4^{3-} + 3\text{H}^+ \rightarrow \text{AlPO}_4 + 3\text{H}_2\text{O}$

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds Cl^- aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver Cl^- ions regardless of the source of these ions these molecular and

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the complete ionic equation for this reaction is as follows equation 4 $13 \text{Ag}^+ \text{aq} + 2 \text{F}^- \text{aq} + 2 \text{NH}_4^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{s} + 2 \text{NH}_4^+ \text{aq} + 2 \text{F}^- \text{aq}$ because two $\text{NH}_4^+ \text{aq}$

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