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Student Exploration: Calorimetry Lab

Vocabulary: calorie, calorimeter, joule, specific heat capacity

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

- The Latin word color means "heat," and meter comes from the Greek word meaning "to measure." What do you think
 a calorimeter does? Measures heat energy in an object
- 2. Where have you heard the word calorie before? What do you think a calorie is? A measure of heat energy

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up. Calorimeters can be used to find a substance's specific heat capacity. You will use the Calorimetry Lab Gizmo to determine the specific heat capacities of various substances.

- On the SIMULATION pane, select Copper. Use the slider to set its Mass to 200 g. Set the Water mass to 200 g. Check that the Water temp is set to 30.0 °C and the copper's Temp is 90 °C. Select the GRAPH tab, and click Play.
- A. What was the Final temperature of the copper and the water? 34.96 oC
- 8. How much did the temperature of the copper change? -55.04 oC
- C. How much did the temperature of the water change? 4.96 oC
- 2. Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain. Water, because the temperature of the water did not change that much. It would have taken more energy to raise the temperature of the water any higher than it already is.

Activity A:

Heat transfer

Get the Gizmo ready:

Click Reset.

Question: What factors determine how heat energy transfers between objects?

Predict: In the Gizmo warm-up, you saw how 200 g of 90 °C copper transfers heat to 200 g

of 30.0 °C water.

- A. How do you think increasing the water's mass would affect the final temperature? It would lower the temperature of the water a little bit
- 8. How do you think decreasing the copper's mass would affect the final temperature? It would increase the temperature of the water a little bit
- C. How do you think increasing or decreasing the copper's initial temperature would affect the final temperature? There would be a slight raise in temperature

Student Exploration Calorimetry Answer Key

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Table of Contents Student Exploration Calorimetry Answer Key

- 1. Understanding the eBook Student Exploration Calorimetry Answer Key
 - The Rise of Digital Reading Student Exploration Calorimetry Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Student Exploration Calorimetry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Student Exploration Calorimetry Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Student Exploration Calorimetry Answer Key
 - Personalized Recommendations
 - Student Exploration Calorimetry Answer Key User Reviews and Ratings
 - Student Exploration Calorimetry Answer Key and Bestseller Lists

- 5. Accessing Student Exploration Calorimetry Answer Key Free and Paid eBooks
 - Student Exploration Calorimetry Answer Key Public Domain eBooks
 - Student Exploration Calorimetry Answer Key eBook Subscription Services
 - Student Exploration Calorimetry Answer Key Budget-Friendly Options
- 6. Navigating Student Exploration Calorimetry Answer Key eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Student Exploration Calorimetry Answer Key Compatibility with Devices
 - Student Exploration Calorimetry Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Student Exploration Calorimetry Answer Key
 - Highlighting and Note-Taking Student Exploration Calorimetry Answer Key
 - Interactive Elements Student Exploration Calorimetry Answer Key
- 8. Staying Engaged with Student Exploration Calorimetry Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Student Exploration Calorimetry Answer Key
- 9. Balancing eBooks and Physical Books Student Exploration Calorimetry Answer Key
 - Benefits of a Digital Library
 - \circ Creating a Diverse Reading Collection Student Exploration Calorimetry Answer Key
- 10. Overcoming Reading Challenges
 - o Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Student Exploration Calorimetry Answer Key
 - Setting Reading Goals Student Exploration Calorimetry Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Student Exploration Calorimetry Answer Key
 - Fact-Checking eBook Content of Student Exploration Calorimetry Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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