

Student Exploration: Bohr Model: Introduction

When light passes through a gas, certain wavelengths of the light are absorbed. The result is a unique **absorption spectrum**. Two examples are shown below.



1. What colors of light are absorbed by hydrogen gas?
violet, blue, green, and red
2. What colors of light are absorbed by helium gas?
violet, blue, green, yellow, and red

Gizmo Warm-up

In 1913, Niels Bohr proposed that the unique spectral lines created by different elements were related to the way electrons were arranged around the nucleus. The *Bohr Model: Introduction* Gizmo explores this connection.

The **laser** shown in the Gizmo can emit **photons**, or particles of light, at a variety of wavelengths. The energy of a photon, measured in **electron volts (eV)**, is inversely proportional to its wavelength. Photons that pass through the gas are detected by the photon detector at the right.



1. With the **Energy (eV)** set to 1 eV, click **Fire**. Did the photon go straight through the gas in the tube, or was it absorbed by the gas?
the photon went straight through the gas tube.
2. Set the **Energy (eV)** to 4 eV, and click **Fire**. What happened this time?
the photon was absorbed by the gas.

Activity A: Absorption spectra

Get the Gizmo ready:

- On the **SIMULATION** pane, select **Lamp**.
- Check that **Gas C** is selected.



Introduction: The smaller the wavelength of a photon, the greater its energy. Humans can see photons with wavelengths between 700 nanometers (red) and 400 nanometers (violet), corresponding to energies between 1.8 and 3.1 eV.

Question: What does the absorption spectrum of an element indicate about its electron configuration?

Student Exploration Bohr Model Introduction Answer Key

G Orfield



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