Acids and Bases Strength

Write True (T) or False (F).

- Stronger acids have weaker conjugate bases and weaker acids have stronger conjugate bases. ____
- Stronger bases have stronger conjugate acids and weaker bases have weaker conjugate acids. ____

Write equations that show how the following reactions occur in water.

- F acts as a base.
- ii. HNO₂ acts as an acid.
- iii. Fe(H₂O)₆3* acts as an acid.
- iv. HCO₃ acts as a base.
- v. HCO₃- acts as an acid.
- vi. Al(H₂O)₄(OH)²⁺ acts as a base.

What volume of 0.45 M LiOH would be needed to neutralize 60.0 mL of 0.15 M HI?

Calculate the pH values of the following solutions of strong acids and bases.

0.001 M HCl

0.76 M KOH

iii. 2.8 x 10-4 M Ba(OH)₂

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J Spring

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