

Atomic Theory Worksheet

Name: _____

John Dalton (1766 – 1844):

John Dalton was an English chemist. His ideas form the atomic theory of matter. Here are his ideas.

Atomic Theory



- All elements are composed (made up) of atoms. It is impossible to divide or destroy an atom.
- All atoms of the same elements are alike. (One atom of oxygen is like another atom of oxygen.)
- Atoms of different elements are different. (An atom of oxygen is different from an atom of hydrogen.)
- Atoms of different elements combine to form a compound. These atoms must be in definite whole number ratios. For example, water is a compound made up of 2 atoms of hydrogen and 1 atom of oxygen (a ratio of 2:1). Three atoms of hydrogen and 2 atoms of oxygen cannot combine to make water.



1803

1. What is the name of John Dalton's theory? _____
2. What are elements made of? _____
3. An atom of hydrogen and an atom of carbon are _____.
4. What are compounds made of? _____
5. The ratio of atoms in HCl is: a) 1:3 b) 2:1 c) 1:1

J. J. Thompson (Late 1800s):

J. J. Thompson was an English scientist. He discovered the electron when he was experimenting with gas discharge tubes. He noticed a movement in a tube. He called the movement cathode rays. The rays moved from the negative end of the tube to the positive end. He realized that the rays were made of negatively charged particles – electrons.

Atomic Structure Model



J. J. Thompson



1904

1. What did J.J. Thompson discover? _____
2. What is the charge of an electron? _____
3. What are cathode rays made of? _____
4. Why do electrons move from the negative end of the tube to the positive end?

5. What was Thompson working with when he discovered the cathode rays?

NUCLEAR MODEL



ERNEST RUTHERFORD



1911

Lord Ernest Rutherford (1871 – 1937):

Ernest Rutherford conducted a famous experiment called the gold foil experiment. He used a thin sheet of gold foil. He also used special equipment to shoot alpha particles (positively charged particles) at the gold foil. Most particles passed straight through the foil like the foil was not there. Some particles went straight back or were deflected (went in another direction) as if they had hit something. The experiment shows:

- Atoms are made of a small positive nucleus; positive nucleus repels (pushes away) positive alpha particles
- Atoms are mostly empty space

1. What is the charge of an alpha particle? _____
2. How did he know that an atom was mostly empty space?

3. What happened to the alpha particles as they hit the gold foil?

4. How did he know that the nucleus was positively charged?

Student 1 Answers Atomic Theory

Franklin Bache



Student 1 Answers Atomic Theory:

Exploration of the Structure of Atom Chandan Sengupta, First Publication April 2019 Revised Publication October 2022 Second Revised Edition July 2023 Third Revised Edition December 2024 Published by Chandan Senguta Printed by IECIT Printing and ublication Services Resource Centre Arabinda Nagar Bankura 722101 W B India Topics Covered Atoms and Molecules Structure of Atom Bonding Mechanism and Chemical Reactions Mechanism of Bonding This book is suitable for students of Class 9 to Class 11 Students aspiring for Pre Medical Entrance Examination can also get adequate support Additional Hard Copies can also be obtained from Chandan Sukumar Sengupta Arabinda Nagar Bankura 722101 WB Write to Us for more materials

Chemistry for Degree Students B.Sc. (Honours) Semester I Madan R.L.,2022 This textbook has been designed to meet the needs of B Sc Honours First Semester students of Chemistry as per the UGC Choice Based Credit System CBCS Maintaining the traditional approach to the subject this textbook lucidly explains the basics of Inorganic and Physical Chemistry Important topics such as atomic structure periodicity of elements chemical bonding and oxidation reduction reactions gaseous state liquid state solid state and ionic equilibrium are aptly discussed to give an overview of inorganic and physical chemistry Laboratory work has also been included to help students achieve solid conceptual understanding and learn experimental procedures

Student's Guide to Fundamentals of Chemistry Jo A. Beran,2013-10-22 Student s Guide to Fundamentals of Chemistry Fourth Edition provides an introduction to the basic chemical principles This book deals with various approaches to chemical principles and problem solving in chemistry Organized into 25 chapters this edition begins with an overview of how to define and recognize the more common names and symbols in chemistry This text then discusses the historical development of the concept of atom as well as the historical determination of atomic weights for the elements Other chapters consider how to calculate the molecular weight of a compound from its formula This book discusses as well the characteristics of a photon in terms of its particle like properties and defines the wavelength frequency and speed of light The final chapter deals with the fundamental components of air and the classification of materials formed in natural waters This book is a valuable resource for chemistry students lecturers and instructors

Chemistry for B.Sc. Students - Semester IV: Quantum Mechanics and Analytical Techniques (NEP-UP) RL Madan, This textbook has been conceptualized for B Sc Fourth Semester students of Chemistry as per the latest curriculum on the directives of NEP 2020 for the Universities of UP Maintaining the traditional approach to the subject this textbook comprehensively covers two papers namely Theory and Practical Parts Quantum Mechanics and Analytical Techniques for Theory Part and Instrumental Analysis for Practical Part The topics covered in the book are Atomic Structure Elementary Quantum Mechanics Molecular Spectroscopy Rotational Vibrational Roman and Electronic Spectra UV Visible Spectroscopy Infrared Spectroscopy ¹H NMR Spectroscopy PMR Introduction to Mass Spectrometry Separation Techniques in the Theory Part The Practical Part covering Molecular Weight Determination Spectrophotometry Spectroscopy and Chromatographic Separations have been presented

systematically to help students in achieving solid conceptual understanding and learn experimental procedures

Student's Solutions Manual to Accompany Atkins' Physical Chemistry C. A. Trapp, M. P. Cady, Carmen Giunta, 2010

This solutions manual provides the authors detailed solutions to exercises and problems in physical chemistry It comprises solutions to exercises at the end of each chapter and solutions to numerical theoretical and additional problems

Foundations of College Chemistry Morris Hein, Susan Arena, 2013-01-01 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years Foundations of College Chemistry Alternate 14th Edition has helped readers master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis

Catalogue of the Reference Library Birmingham Free Libraries. Reference Department, John Davis Mullins, 1890 *Minds-on Physics: Fundamental forces & fields* William J. Leonard, 1999 There is one Teacher's Guide which corresponds with each Student Activities Book and consists of two parts Answers and Instructional Aids for Teachers and Answer Sheets The Answers and Instructional Aids for Teachers provides advice for how to optimize the effectiveness of the activities as well as brief explanations and comments on each question in the student activities The Answer Sheets may be duplicated and distributed to students as desired Use of the Answer Sheets is particularly recommended for activities requiring a lot of graphing or drawing *A System of Chemistry for the Use of Students of Medicine* Franklin Bache, 1819

Resources for Teaching Middle School Science Smithsonian Institution, National Academy of Engineering, National Science Resources Center of the National Academy of Sciences, Institute of Medicine, 1998-04-30 With age appropriate inquiry centered curriculum materials and sound teaching practices middle school science can capture the interest and energy of adolescent students and expand their understanding of the world around them Resources for Teaching Middle School Science developed by the National Science Resources Center NSRC is a valuable tool for identifying and selecting effective science curriculum materials that will engage students in grades 6 through 8 The volume describes more than 400 curriculum titles that are aligned with the National Science Education Standards This completely new guide follows on the success of Resources for Teaching Elementary School Science the first in the NSRC series of annotated guides to hands on inquiry centered curriculum materials and other resources for science teachers The curriculum materials in the new guide are grouped in five chapters by scientific area Physical Science Life Science Environmental Science Earth and Space Science and Multidisciplinary and Applied Science They are also grouped by type core materials supplementary units and science activity books Each annotation of curriculum material includes a recommended grade level a description of the activities involved and of what students can be expected to learn a list of accompanying materials a reading level and ordering information The curriculum

materials included in this book were selected by panels of teachers and scientists using evaluation criteria developed for the guide. The criteria reflect and incorporate goals and principles of the National Science Education Standards. The annotations designate the specific content standards on which these curriculum pieces focus. In addition to the curriculum chapters, the guide contains six chapters of diverse resources that are directly relevant to middle school science. Among these is a chapter on educational software and multimedia programs, chapters on books about science and teaching, directories and guides to science trade books and periodicals for teachers and students. Another section features institutional resources. One chapter lists about 600 science centers, museums, and zoos where teachers can take middle school students for interactive science experiences. Another chapter describes nearly 140 professional associations and U.S. government agencies that offer resources and assistance. Authoritative, extensive, and thoroughly indexed, and the only guide of its kind, *Resources for Teaching Middle School Science* will be the most used book on the shelf for science teachers, school administrators, teacher trainers, science curriculum specialists, advocates of hands-on science teaching, and concerned parents.

Adolescent Literature as a Complement to the Content Areas Paula Greathouse, Joan F. Kaywell, Brooke Eisenbach, 2017-08-10. This text offers 6th-12th grade educators guided instructional approaches for including young adult (YA) literature in science and math classes in order to promote literacy development while learning content. Chapters are co-authored, pairing content experts with literacy experts to ensure that both content and literacy standards are met in each approach. Each chapter spotlights the reading of one YA novel and offers pre-reading and after-reading activities that guide students to a deeper understanding of the content while increasing their literacy practices. While each chapter focuses on a specific content topic, readers will discover the many opportunities reading YA literature in the content area has in encouraging cross-disciplinary study.

Chemistry: for B.Sc. Students Semester II (NEP-2020 KSHEC) R. L. Madan, This textbook has been conceptualised for universities of Karnataka as per the recommended National Education Policy (NEP) 2020 to meet the needs of B.Sc. students of Chemistry. Maintaining the traditional approach to the subject, this textbook lucidly explains the basics of Inorganic and Physical Chemistry. Important topics such as atomic structure, periodicity of elements, kinetic theory of gases, liquid state, liquid crystals, and types of solids are aptly discussed to give an overview of inorganic and physical chemistry. Laboratory work (33 experiments) has also been included to help students achieve sound conceptual understanding and learn experimental procedures.

Chemical Interactions, 2005. **ACT Math & Science Prep** Kaplan Test Prep, 2016-03-01. Kaplan Test Prep is the Official Partner for Live Online Prep for the ACT. For more information, visit [kaptest.com](https://www.kaptest.com).

onlineprep. The complete ACT test prep tool that contains comprehensive practice and tips for students who want to sharpen their Math and Science skills and score higher on test day. In 2015, approximately 1.9 million high school students took the ACT. Despite the popularity of the ACT, 58 percent of students are not reaching the readiness benchmark for Math. ACT Math Science Prep will help you prepare for these challenging sections of the ACT. This comprehensive tool contains

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