

Student Exploration: Calorimetry Lab Vocabulary: calorie, calorimeter, joule, specific heat capacity

Calorimetry Lab

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. The Latin word *calor* means “heat,” and *meter* comes from the Greek word meaning “to measure.” What do you think a **calorimeter** does? **< Measures heat >**
2. Where have you heard the word **calorie** before? What do you think a calorie is? **< unit of measurement for heat. >**

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance’s **specific heat capacity**. You will use the *Calorimetry Lab* Gizmo™ to determine the specific heat capacities of various substances.



1. On the SIMULATION pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper’s **Temp** is 90 °C. Select the GRAPH tab, and click **Play** (▶).
 - A. What was the **Final temperature** of the copper and the water? **< 30.54 >**
 - B. How much did the temperature of the copper change? **< 60 >**
 - C. How much did the temperature of the water change? **< 5 >**
2. Specific heat capacity can be described as a substance’s resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.
< Copper because it’s a better heat conductor than water. >

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