Straighterline Chemistry Final- deck 3

Which is more soluble in an acidic solution than in pure water? Select one:

- a. Cull
- b. PbCl2
- c. Ca3(PO4)2
- d. NaNO3
- e. NaBr ANSc. Ca3(PO4)2

What is the name of the principle of selective precipitation used to identify the types of ions present in a solution?

Select one:

- Ionization
- b. Selective ion precipitation
- c. Selective ion typing
- d. Limited precipitation
- e. Qualitative analysis ANSe. Qualitative analysis

A 10.0-mL sample of 0.75 M CH3CH2COOH is titrated with 0.30 M NaOH. What is the pH of the solution after 22.0 mL of NaOH have been added to the acid?

 $[Ka(CH3CH2COOH) = 1.3 \times 10-5]$

Select one:

- a. 5.75
- b. 4.94
- c 4.83
- d. 4.02
- e. 3.95 ANSa. 5.75

Methyl red is a common acid-base indicator. It has a Ka equal to 6.3 × 10-6. Its un-ionized form is red and its anionic form is yellow. What color would a methyl red solution have at pH = 7.8? Select one:

- a. green
- b. red
- c. blue
- d. yellow
- e. violet ANSd. vellow

A 25.0-mL sample of 0.10 M C2H5NH2 (ethylamine) is titrated with 0.15 M HCl. What is the pH of the solution after 9.00 mL of acid have been added to the amine? [Kb(C2H5NH2) = 6.5 × 10-4] Select one:

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