

Modeled after AP 2013 #21:



In the reaction represented above, what is the total number of moles of reactants consumed when 2.00 mole of CO_2g is produced?

- A. 0.50-mol
- B. 1.50-mol
- C. 4.50-mol
- D. 9.00-mol



In the reaction represented above, what is the total number of moles of reactants consumed when 2.00 mole of CO_2g is produced?

- A. 0.50-mol
- B. 1.00-mol
- C. 3.50-mol
- D. 7.00-mol



In the reaction represented above, what is the total number of moles of products made when 3.50 mole of O_2g is consumed?

- A. 2.00-mol
- B. 3.00-mol
- C. 5.00-mol
- D. 10.0-mol



In the reaction represented above, what is the total number of moles of products made when 1.50 mole of O_2g is consumed?

- A. 1.00-mol
- B. 2.00-mol
- C. 4.00-mol
- D. 8.00-mol

Modeled after AP 2014g #55:



A 7.0 mol sample of NH_3g and a 10. mol sample of O_2g are placed in a previously evacuated container and react according to the equation above. How many moles of NOg are produced when one of the reactants has been totally consumed?

- A. 2.5 mol
- B. 7.0 mol
- C. 8.0 mol
- D. 18 mol



Samples containing 4.0 mol $\text{C}_2\text{H}_6\text{g}$ and 5.0 mol H_2g are placed in a previously evacuated container and allowed to react according to the equation above. When one of the reactants is completely consumed, how much product is formed?

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Clifford Lane



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