UNIT VII Review Stoichiometry

Stoichiometry Calculations involving MASS, MOLES, VOLUME, and MOLECULES

- Consider the reaction 4NH₃ +5O₂ → 6H₂O + 4NO
 - a. What mass of NO is produced when 5.00 mol of NH3 are reacted with excess O₂? | 50 9 NO
 - b. What volume of O2 at STP is required to produce 90.0L of NO? WS . O.
 - c. How many molecules of H₂O are formed when 100.0g of NH₁ react?
- Consider the reaction SiCl₄ + 2H₂ → Si + 4HCl
 - a. How many moles of SiCl, are required to produce 4.25 x 10¹² molecules of HCl
 - b. If 25.0 mL of H2 (g) at STP react, calculate the mass of Si produced 0 0157 5
 - c. If 0.892 kg of SiCl. react, calculate the number of molecules of HCl produced. 21.0 mg/ HCl
- Silver reacts with nitric acid, HNO₃ to form nitrogen monoxide, NO, silver nitrate, AgNS₃ and water
 - a. How man spens of mitric acid are required to react with 5 00g of silver?
 - b. How sharry grams of silver nitrate will be formed?
 - c. How many moles of nitrogen monomide will be produced?

Stoichiometry using MOLAR CONCENTRATION

- Consider the following reaction CaCO₃ + 2HCl → CaCl₂ + CO₂ + H₂O
 - a. What volume of HCl having a concentration of 0.0010M is neutralized by a 0.850 g portion of CaCO₃? 17 OLHCI
 - b. What volume of CO2 at STP1s produced it 1 25L of 0 0055 M HCl reacts with an excess of CaCO3? O 077L (01)
 - c. How many moles of 0 200 M HCl is reacted to produce 3 00g of CaCl₂?
- If 17.5 g of zinc are reacted with phosphoric acid, H₃PO₆, then zinc phosphate, Zn₃(PO₆)₃, and hydrogen are produced.
 - a How many moles of phosphoric acid are required 70 178 mol 45 Pou
 - b. If the phosphoric acid is a 3 00M solution, how many litres of the solution are needed?
 - c. What mass of zinc phosphate will be produced?

(84 49 Zng (104)2)

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