ssessment #3 Review	Name:
ou must show all work for	credit.
ie Date:	(No late packets will be accepted)
I. Balance Equation	ns.
Balance the equations	F II True
<ol> <li>Mg + ∠ H<sub>2</sub>O → M;</li> </ol>	$g(OH)_2 + H_2$
<ol> <li>Mg + ZNaBr → N</li> </ol>	MgBr <sub>2</sub> + ZNa
3. C <sub>3</sub> H <sub>8</sub> + 5 O <sub>2</sub> →3 C	
4. 4 P + 5 O <sub>2</sub> → 2 P <sub>2</sub> O <sub>5</sub>	1 1 1 1 CO 1 7 Ec/OH)
5. Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub> + 6 KOH	I →3 K₂SO₄ + 2 Fe(OH)₃
<ol> <li>Fe<sub>2</sub>O<sub>3</sub> + 3 H<sub>2</sub> → F</li> </ol>	7
7 How many moles of	hydrogen gas react with 1 mole of Fe₂O₃ in reaction #6?
7. 11011 1112117	
rite the balanced equation	n for problems # 8 and 9:
<ol><li>Hydrogen gas and n</li></ol>	nitrogen monoxide combine to form water and nitrogen gas
C Man and make a	phate and calcium chloride are mixed they form calcium phosphate and
sodium chloride	2 Nas 704 13 Cacle - Cas (PON) = + 8 Na C
Sodium chloride	< Nastry 13 cacks - castronie actions
10 Using the following b	balanced equation, prove (or explain) the Law of Conservation of Mass.
$N_2 + 3H_2 \rightarrow 2NH_3$	
	Z N = lety Z = 1 ; [w]
	su no plants created or lost
The second second second second	equation shown in problem #10, explain what the coefficients represent.
11. Using the balanced	equation shown in problem #10, explain what the coemoleria represents
12. Using the balanced	equation shown in problem #10, list the products of the reaction.
	UKZ
13. Using the balanced	equation shown in problem #10, list the reactants of the reaction.
	1 12
II. Types of Reacti	ions
irections: Identify the type	e of reaction.
= synthesis	SD = single displacement C= combustion DD = double displacement
= decomposition	DD = double displacement
14. HCI + MgS	$O_4 \rightarrow MgCl_2 + H_2SO_4$
15. Cu₂(CO₃)₃ →	Cu + CO <sub>2</sub> + O <sub>2</sub>
16 C <sub>2</sub> H <sub>6</sub> + O <sub>2</sub>	→ CO <sub>2</sub> + H <sub>2</sub> O
	→ CO <sub>2</sub> + H <sub>2</sub> O
16. $C_2H_6 + O_2$	$\rightarrow$ CO <sub>2</sub> + H <sub>2</sub> O P <sub>2</sub> O <sub>5</sub>
16. $C_2H_6 + O_2$	$\rightarrow$ CO <sub>2</sub> + H <sub>2</sub> O P <sub>2</sub> O <sub>5</sub> CaCl <sub>2</sub> $\rightarrow$ NaCl + Ca(NO <sub>3</sub> ) <sub>2</sub>

# **Unit 6 Test Chemistry**

Ajnish Kumar Gupta, Bharti Gupta

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