

CHEMISTRY SAMPLE TEST 1

$$pV = nRT \rightarrow n = \frac{pV}{RT}$$

$$n = 5,003$$

1. A sample of N_2 gas occupies 4.48 L volume under standard conditions (1 atm, $0^\circ C$). What is the mass of the sample? The atomic mass of N is 14 amu.

A: 5.6 g B: 56 g C: 2.8 g D: 28 g E: 22.4 g

2. Which properties are characteristic for the nonmetals?

- ~~1.~~ high electrical conductivity
~~2.~~ large ionization energy
~~3.~~ high electronegativity
~~4.~~ low electron affinity

~~A: 1,2~~

B: 2,3

C: 3,4

D: 2,4

~~E: 1,2,3~~

3. Concerning 1.2×10^{24} CO_2 molecules, which statements are true?

The molar mass is 44 g/mol.

1. it is 12 moles.
 2. it occupies 1.2×22.4 L volume under standard conditions.
 3. it has a mass of 88 grams.
 4. it consists of 3.6×10^{24} atoms.

$$\text{mole} = \frac{\text{particle}}{\text{Avogadro's no}}$$

A: 1,3

B: 1,4

C: 2,3

D: 2,4

E: 3,4

4. Which of the following substances contain covalent bonds only?

~~1.~~ $BaCl_2(aq)$ *metal*

2. $CCl_4(l)$ *covalent*

3. $HCl(g)$ *covalent*

• **covalent bond:** two atoms share electrons
 • **ionic bond:** one atom donates electron to the other
 ↑ **affinity**
 1 non-metal + 1 metal

A: 1,2,3,4

B: 1,2,4

C: 2,3

D: 1,3

~~E: 2,3,4~~

5. Which of the following atoms are isotopes of each other?

- ~~X:~~ 11 protons, 11 electrons, 12 neutrons
~~Y:~~ 11 protons, 10 electrons, 12 neutrons
~~Z:~~ 11 protons, 11 electrons, 13 neutrons
~~W:~~ 12 protons, 12 electrons, 12 neutrons

isotopes
 Same proton no.
 Different neutron no.

A: X and Y

B: X and W

C: Y and Z

D: Z and W

~~E: X and Z~~

6. An aqueous solution is prepared by dissolving 6 g NaOH in 250 mL final volume. What is the molar concentration of the solution?

The molar mass of NaOH is 40 g/mol

- A: 6.4 mol/L
 B: 1.6 mol/L
 C: 16 mol/L
 D: 0.16 mol/L
 E: 64 mol/L

$$C_M = \frac{n}{V}$$

$$\frac{6 \text{ mol}}{40} = 0.15 \text{ mol} \quad \frac{250}{1000} \text{ L}$$

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Raffaela Di Napoli



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