

Unit 10 Stoichiometry Test 1: Short Answers

Answer all questions including units and balanced equations where appropriate.

1. How many moles of aluminium are produced when completely with 2.5 mol of FeCl_3 ?

$$2\text{Al} + 3\text{FeCl}_2 \rightarrow 3\text{Fe} + \text{Al}_2\text{Cl}_6$$
2. How many grams of FeCl_3 are produced when 2.5 moles of aluminium reacts completely with FeCl_2 ?

$$2\text{Al} + 3\text{FeCl}_2 \rightarrow 3\text{Fe} + \text{Al}_2\text{Cl}_6$$
3. The reaction below shows that decomposition of Hg_2Cl_2 occurs. How many grams of mercury are produced when 20.7g Hg_2Cl_2 is heated?

$$\text{Hg}_2\text{Cl}_2 \rightarrow 2\text{Hg} + \text{Cl}_2$$
4. An element that has been calculated to produce 2.0g of HgCl_2 actually produces 2.2g of HgCl_2 . What is the percent yield?
5. How many moles of hydrogen gas are produced when 20.0g of zinc reacts with 20.0g of hydrochloric acid according to this equation.

$$\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$$

6. Percent Yield is

7. Amount of 20.0g produced is

8. How actually when 20.0g of Hg_2Cl_2 what is the percent yield?

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JE Gale



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