

10**CHEMICAL QUANTITIES****Chapter Test A****A. Matching**

Match each description in Column B to the correct term in Column A. Write the letter of the correct description on the line.

Column A

- _____ 1. standard temperature and pressure
- _____ 2. molar volume
- _____ 3. representative particle
- _____ 4. molar mass
- _____ 5. atomic mass
- _____ 6. Avogadro's number
- _____ 7. mole (mol)
- _____ 8. empirical formula
- _____ 9. percent composition
- _____ 10. molar mass

Column B

- a. the volume occupied by a mole of any gas at STP (22.4 L)
- b. the mass in amu of an atom of an element
- c. the number of representative particles in a mole of a substance
- d. the percent by mass of each element in a compound
- e. the mass of a mole of a substance
- f. the formula that gives the lowest whole-number ratio of the elements in a compound
- g. an atom, a formula unit, or a molecule
- h. the mass in grams of a mole of any pure substance
- i. 6.02×10^{23} representative particles of a substance
- j. 0°C and 101.3 kPa

B. Multiple Choice

Choose the best answer and write its letter on the line.

- _____ 11. How many hydrogen atoms are in 4 molecules of isopropyl alcohol, $\text{C}_3\text{H}_8\text{O}$?
 - a. $4 \times (6.02 \times 10^{23})$
 - b. 4
 - c. 32
 - d. $32 \times (6.02 \times 10^{23})$
- _____ 12. The mass of a mole of NaCl is its:
 - a. molar mass.
 - b. empirical formula.
 - c. gram atomic mass.
 - d. atomic mass.

Unit 6 Chapter 10 Chemical Quantities Test Review

Rachel S Tattersall



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