

Name \_\_\_\_\_  
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## Unit 6: Stoichiometry – Intensive Review

### Formula Mass & Percent Composition

1. The sum of the atomic masses of the atoms in one molecule of  $C_4H_{10}$  is called the

- (1) formula mass                      (3) percent abundance  
(2) isotopic mass                    (4) percent composition

2. The gram-formula mass of a compound is 48 grams. The mass of 1.0 mole of this compound is

- (1) 1.0 g                      (2) 4.8 g                      (3) 48 g                      (4) 480 g

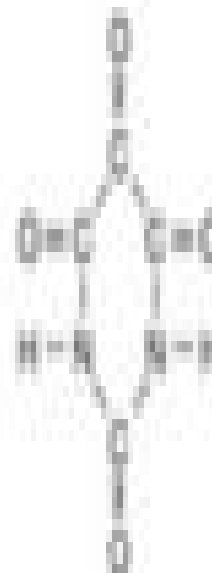
3. What is the gram-formula mass of  $(NH_4)_3PO_4$ ?

- (1) 112 g/mol                      (2) 121 g/mol                      (3) 149 g/mol                      (4) 242 g/mol

4. What is the gram-formula mass of  $Ca_3(PO_4)_2$ ?

- (1) 248 g/mol                      (2) 263 g/mol                      (3) 279 g/mol                      (4) 310 g/mol

14. Given the formula for a compound,



Which molecular and empirical formulas represent this compound?

- (1)  $C_2H_4O_2$  and  $CH_2O$                       (3)  $C_2H_4N_2O_2$  and  $CH_2NO$   
(2)  $C_2H_4O_2$  and  $C_2H_4O$                       (4)  $C_2H_4N_2O_2$  and  $C_2H_4NO$

15. A compound has a molar mass of 90. grams per mole and the empirical formula  $CH_3O$ . What is the molecular formula of the compound?

- (1)  $CH_3O$                       (2)  $C_2H_6O_2$                       (3)  $C_3H_9O_3$                       (4)  $C_4H_{12}O_4$

16. The structural formula for methanol is  $CH_3$  and the molecular

# Unit 6 Answer For Chemistry Regents Review

**Michal Rosen-Zvi**



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