

# A Level Required Practical 3 - Rate of Reaction

## Theory

The time taken for a cross placed under the reaction mixture to disappear from the cloudiness of the sulfur.



Measure rate of reaction as  $1/\text{time}$ .

## Apparatus

- Thermometer
- Large Beaker to be used as water bath
- 2 Boiling tubes
- $0.05 \text{ mol dm}^{-3}$  of sodium thiosulfate solution
- $1.0 \text{ mol dm}^{-3}$  of hydrochloric acid
- Measuring cylinder
- Graduated pipette
- Stopwatch
- Graph paper

## Method

1. Using a measuring cylinder, measure  $10 \text{ cm}^3$  of the hydrochloric acid to an 'acid' tube.
  - a.  $10 \text{ cm}^3$  of sodium thiosulfate into another tube and place this tube over the cross and also put in a thermometer.
2. Note the starting temperature (room temperature) then add  $1 \text{ cm}^3$  of the acid to the sodium thiosulfate and start the timer.
3. Look down at the tube from above and record the time it takes for the cross to disappear. Also record the temperature of the reaction mixture at the end, pour the contents into the sodium carbonate solution.
4. Measure another  $10.0 \text{ cm}^3$  of sodium thiosulfate solution into a clean tube and place it in a water bath (large beaker) (no hotter than  $55^\circ\text{C}$  otherwise too much  $\text{SO}_2$  produced) to warm up for 3 minutes.
5. Repeat the steps 2 and 3.
6. Repeat these steps for 5 different temperatures.

## Variables

- Independent variables

# Unit 16 Reaction Rates Notes Answers

**Marco Cascella**



## Unit 16 Reaction Rates Notes Answers:

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