Name:		Date:	
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Unit 7 Chemistry - Nail Lab

Purpose After you have completed the lab, think about what was the overall point for performing the lab. What were we trying to accomplish?

Procedure

Day 1

- Mass a clean & labeled mL beaker.
- Put between _____ g of copper (II) chloride in the beaker. Record the mass of the beaker and the copper (II) chloride.
- Add mL of distilled water to the beaker. Stir to dissolve the solid.
- Obtain the mass of ______ iron nails together to ± 0.01g.
- Place the nails in the copper (II) chloride solution. Observe the reaction; record your observations. Place the labeled beaker in the place designated by your teacher.

when you decant. After a 2nd rinse with distilled water, rinse the copper with

Day 2

- 7. _____ the solution from the 250 mL beaker. Rinse the precipitate
 with about mL of distilled water. Try to lose as little of the solid copper as you can

. Rinse one last time with distilled water. Then place the labeled beaker in the drying oven.

8. Mass the dry nails, and then discard the nails.

Day 3

Mass the beaker + dry copper. Discard the copper in the place designated by your teacher. Wash
your beaker and let dry.

Data:

Mass 250 mL beaker	
Mass 250 mL beaker + copper (II) chloride	
Mass nails before reaction	
Mass nails after reaction	

Unit Nail Lab Chemistry Conclusion Answers

J Ma

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