

Unit 2 Worksheet: St. Michael's Concentration

Part A: Dissolving Metal Concentration

1. Cadmium chloride has a molar mass of 208.42 g/mol. Explain what molar concentration tells you about this CdCl_2 solution. What might a chemical use this value?

2. A 20.00 g sample of potassium nitrate is dissolved in enough water to make 200.0 mL of solution. Calculate the molar concentration of this potassium nitrate.

3. Find the molarity of a solution made from 2.50 g of CaCl_2 dissolved in enough water to make 0.100 L.

4. An unknown metal reacts with hydrochloric acid and produces hydrogen gas. Dissolving 0.20 g of metal produces 0.015 moles of hydrogen gas. Calculate the molarity of the metal in the solution.

5. What metal cation would be produced in a reaction of 0.20 g of metal with 0.015 moles of HCl ?

Unit 9 Worksheet 3 Molar Concentration Part B

IM Harris



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