

Unit 8 – Stoichiometry Worksheet 1: Mole relationships

For each of the problems below, use the five steps discussed in class. Answer on a separate piece of paper.

1. Write the balanced chemical equation
2. Identify what is given (with units) and what you want to find (with units) in a BCA table
3. Use the mole ratio from the balanced equation to calculate remaining values in the Table (dimensional analysis style). **Show your work beneath the BCA table.**
4. Complete the BCA Table using calculation results
5. Find the desired quantity in your work. Write and box in that answer with unit, substance.

1. Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8.0 moles of hydrogen sulfide? (The BCA table is done as an example. Make/copy your own on your own piece of paper).

Equation: $\text{___ H}_{2}\text{S}_{(g)} + \text{___ O}_{2(g)} \rightarrow \text{___ SO}_{2(g)} + \text{___ H}_2\text{O}_{(g)}$

Before:

FILL IN THE VALUES ON YOUR OWN PAPER

Change

(SHOW WORK UNDER THE TABLE)

After

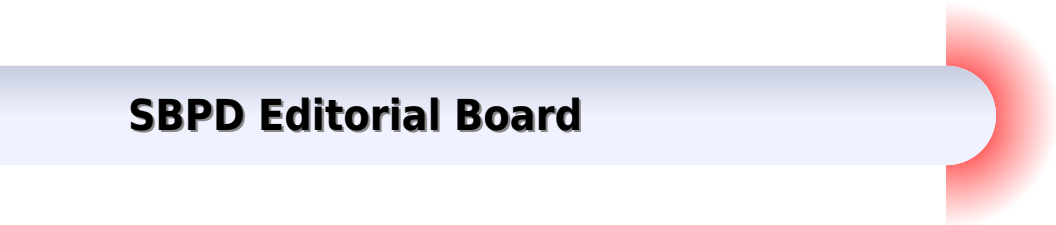
2. Propane, C_3H_8 , burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?
3. Ammonia, NH_3 , for fertilizer is produced by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?
4. The poison gas phosgene, COCl_2 , reacts with water in the lungs to form hydrochloric acid and carbon dioxide. How many moles of hydrochloric acid would be formed by 0.835 moles of phosgene?
5. Nitric acid, HNO_3 , is used in the production of trinitrotoluene, (TNT), an explosive. The equation (not balanced) is



- a) How many moles of nitric acid are required to react with 4.5 moles of toluene? (make your own BCA table)
 - b) How many moles of TNT should be produced?
6. Iron metal and oxygen combine to form magnetite, the magnetic oxide of iron, Fe_3O_4 .
 - a) How many moles of iron can be converted to magnetite by 8.80 moles of pure oxygen? (make your own BCA table)
 - b) How many moles of the magnetic iron oxide would be produced?

Unit 8 Stoichiometry Answers

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Th Thorium R. Benz, A. Naoumidis, D. Brown, 2013-11-11 This volume C 3 as a part of the Gmelin Thorium Handbook Series C describes the thorium nitrogen compounds Included are compounds both of technological importance like the nitrides and the nitrates and those of merely scientific interest such as amides and related compounds However due to the decreasing technical importance of the nuclear thorium fuel cycle especially with the advanced fuels like the nitride ThN in recent years the thorium compounds with nitrogen have been investigated much less extensively than the corresponding uranium compounds In order to have the data for the Th N X systems accumulated in one specific volume the decision was made to publish this volume without incorporating other Th systems ThN is the compound with the lowest N Th ratio In addition to its former nuclear interest due to its thermal and radiation stability it has many very interesting physicochemical properties Thorium nitrate the other well investigated compound is of importance because it is in the form of an adduct with tri n butylphosphate the extracted compound when burnt up thorium fuels are reprocessed Despite the wealth of accumulated data on the chemical and physicochemical properties of the compounds discussed the knowledge of the compounds and of the systems is far from satisfactory it must be deepened and improved in further studies I would like to thank the competent authors for their critical contributions as well as the Gmelin Institute for the excellent cooperation provided especially Prof Dr Fluck and Dr Keim the editor in chief of this volume

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