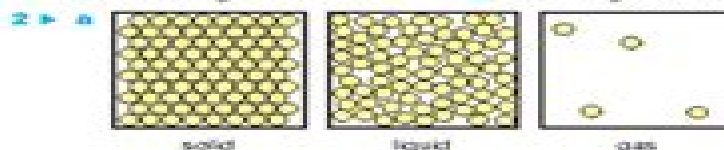


UNIT 1 ANSWERS

CHAPTER 1

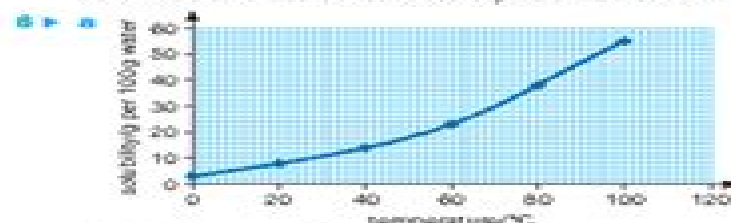
- 1 ▶ a melting b freezing
c subliming/sublimation d subliming/sublimation



Notes: Solids should have regularly packed particles touching. Liquids should have most of the particles touching at least some of their neighbours, but with gaps here and there, and no regularity. Gases should have the particles well spaced.

- b Solids: vibration around a fixed point. Liquids: particles can move around into vacant spaces, but with some difficulty because of the relatively close packing.
c Evaporation: Some faster moving particles break away from the surface of the liquid. Boiling: Attractive forces are broken throughout the liquid to produce bubbles of vapour.
- 3 ▶ a i A – gas; B – liquid; C – solid; D – liquid; E – solid
ii A – gas; B – solid; C – solid; D – liquid; E – solid
iii A – gas; B – liquid; C – solid; D – gas; E – solid
b A, because it is a gas.
c It sublimates and therefore is converted directly from a solid to a gas without going through the liquid stage.
d D – because it has a lower boiling point the forces of attraction between particles will be weaker therefore it will also evaporate more easily than substance B (the only other substance that is a liquid at 25°C).
- 4 ▶ a The ammonia and hydrogen chloride particles have to diffuse through the air in the tube, colliding with air particles all the way.
b i Its particles will move faster.
ii It would take slightly longer for the white ring to form, because the gas particles would be moving more slowly at the lower temperature.
c Ammonia particles are lighter than hydrogen chloride particles and so move faster. The ammonia covers more distance than the hydrogen chloride in the same time.
d i Ammonium bromide.
ii The heavier hydrogen bromide particles would move more slowly than the hydrogen chloride particles, and so the ring would form even closer to the hydrobromic acid end than it was to the hydrochloric acid end. The ring will also take slightly longer to form because of the slower moving particles.

- 5 ▶ Sodium chloride dissolves in water to form a solution. The water is called the **solvent** and the sodium chloride is the **solute**. If the solution is heated to 50°C some of the water **evaporates** until the solution becomes **saturated** and sodium chloride crystals start to form.



- b 94 $\pm$ 1 g per 100 g
The values obtained in this question and in c depend on the line of best fit. In the exam there will always be some tolerance – a range of values will be accepted.
c From the graph, the solubility at 30°C is 1.0 g per 100 g of water.
 $\frac{40}{100} \times 10 = 4$ g
Therefore 4 g of sodium chloride will dissolve.
d i 53 $\pm$ 1 °C
ii The solubility at 17 °C is 7 $\pm$ 1 g per 100 g, therefore 20 – 7 = 13 g must precipitate out of the solution. Answers of 13 $\pm$ 1 g are acceptable.

CHAPTER 2

- 1 ▶
- | Element | Compound | Mixture |
|----------|--------------------|---------------------------|
| hydrogen | magnesium oxide | sea water |
| calcium | copper(II) sulfate | honey |
| | | blood |
| | | mud |
| | | potassium iodide solution |
- 2 ▶ a mixture b mixture c element
d element e compound f compound
- 3 ▶ Substance X is the pure substance – it melts at a fixed temperature. Substance Y is impure – it melts over a range of temperatures.
- 4 ▶ a crystallisation b (simple) distillation
c fractional distillation d chromatography
e filtration
- 5 ▶ For example: Stir with a large enough volume of cold water to dissolve all the sugar. Filter to leave the diamonds on the filter paper. Wash on the filter paper with more water to remove any last traces of sugar solution. Allow to dry.
- 6 ▶ a M b R
c 0.45 $\pm$ 0.01 (measure to the centre of the spot and remember to measure from the base line and not from the bottom of the paper)
d G and T e P

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S Baum



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