

IN CLASS - Electron Configuration Skill Review

Name: _____

Learning Objective: Prepare for writing accurate electron configurations & use the location of the outermost electrons to explain chemical reactivity.

3) What is an electron configuration? Be specific - use an example.

An electron configuration is similar to writing down an "address" for the highest likelihood of locating of an electron. The electron configurations use a code with numbers representing principle energy levels, the letters s, p, d & f to represent sublevels and subscript values to represent the number of electrons in each sublevel.

For example: C (6e-) = $1s^2 2s^2 2p^2$

Write the complete electron configuration for Phosphorus (P).

P (15e-) = $1s^2 2s^2 2p^6 3s^2 3p^3$

2) Write the complete electron configuration of Iodine (I).

I (53e-) = $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^5$

3) Underline the valence electrons in your above electron configuration of Phosphorus and explain the relationship between chemical reactivity, valence electrons and arrangement of the periodic table.

P (15e-) = $1s^2 2s^2 2p^6 \underline{3s^2 3p^3}$ ← P has 5 val e- located in its outermost energy level. This means that P would be located in the 5th group on the Periodic Table and have similar chemical properties to the other elements in group 5.

4) Use the Wave Mechanical Model to explain the location of Phosphorus' 15th electron from the nucleus using principle energy level, sublevel, orbital and spin. (Can use an orbital diagram to enhance your explanation.)

P (15e-) = $1s^2 2s^2 2p^6 3s^2 3p^3$

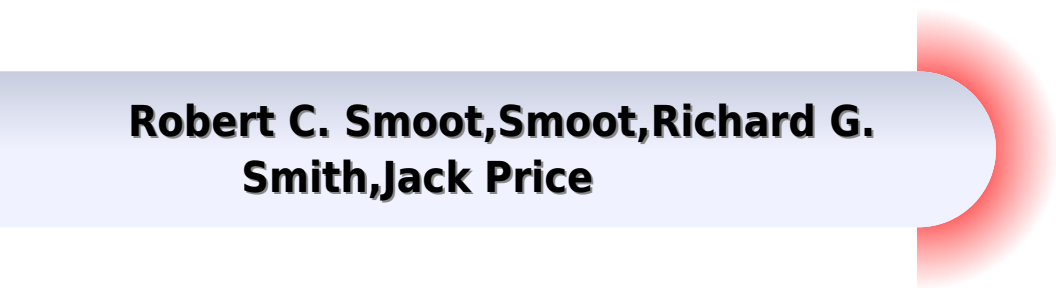
Phosphorus' 15th electron is found in the 3rd principle energy level, in the p sublevel, in its 3rd orbital, spinning by itself.



Unit 6 Electron Configuration And Periodicity Review

Answer Key

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**



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