

19. On the basis of enthalpy and entropy, predict whether each of the following reactions would be spontaneous as written or not at room temperature.



Minimum enthalpy favours (reactants/products) reactants
 Maximum entropy favours (reactants/products) reactants
 Spontaneous as written? (yes/no) Answer no



Minimum enthalpy favours (reactants/products) products
 Maximum entropy favours (reactants/products) products
 Spontaneous as written? (yes/no) Answer yes



Minimum enthalpy favours (reactants/products) reactants
 Maximum entropy favours (reactants/products) products
 Spontaneous as written? (yes/no) Answer yes, but it would

not go to completion.
A state of equilibrium
would be reached.

Unit 4 Section A Supplement Answers Chemistry

Daniel F McAuley



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