

## Unit 1 Chemistry Study Guide

### History of the Atom:

#### Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called **atomos**.

#### Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called **hyle**.

#### Antoine Lavoisier (1774):

- **Law of Conservation of Mass:** Mass cannot be created or destroyed.

#### Joseph Proust (1779):

- **Law of Definite Proportions:** All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

#### John Dalton (1803 – 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first **atomic theory**.
- **Dalton's Atomic Theory:**
  - o All matter is composed of extremely small particles called atoms.
  - o All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
  - o Atoms cannot be subdivided, created, or destroyed.
  - o Atoms of different elements can combine in simple whole number ratios to form chemical compounds.
  - o In chemical reactions, atoms are combined, separated or rearranged.
- **Law of Multiple Proportions:** Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

#### Billiard Ball Model:



#### J.J. Thompson (1906):

- **Cathode Ray Experiment –**
  - o Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

# Unit 1 Chemistry Study Guide Answers

**Léo Pomar**



## **Unit 1 Chemistry Study Guide Answers:**

## Enjoying the Track of Term: An Psychological Symphony within **Unit 1 Chemistry Study Guide Answers**

In a global consumed by monitors and the ceaseless chatter of instant connection, the melodic elegance and psychological symphony created by the published term frequently diminish into the back ground, eclipsed by the persistent noise and distractions that permeate our lives. However, located within the pages of **Unit 1 Chemistry Study Guide Answers** a charming fictional treasure filled with raw feelings, lies an immersive symphony waiting to be embraced. Constructed by an elegant musician of language, that fascinating masterpiece conducts visitors on a mental trip, well unraveling the hidden melodies and profound influence resonating within each carefully constructed phrase. Within the depths with this moving review, we shall examine the book is central harmonies, analyze their enthralling publishing fashion, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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