AP Chemistry - Unit 9 Practice Test (Thermochemistry and Electrochemistry - Chapters 19 and 20)

The value of ΔG° at 141.0 °C for the formation of phosphorous trichloride from its constituent
 elements

 $P_2(g) + 3Cl_2(g) \rightarrow 2PCl_3(g)$

is _____ kJ/mol. At 25.0°C for this reaction, ΔH^o is -720.5 kJ/mol, ΔG^o is -642.9 kJ/mol, and ΔS^o is -263.7 J/K.

A) -683.3

B) 3.65×10^4

C) -829.7

D) -612.3

E) 1.08×10^5

2) How many grams of copper will be plated out by a current of 2.3 A applied for 25 minutes to a 0.50
2) ______
M solution of copper(II) sulfate?

A) 1.8 × 10-2

B) 2.2

C) 0.019

D) 0.036

E) 1.1

The standard cell potential (E*) of a voltaic cell constructed using the cell reaction below is 0.76 V:

$$Zn(s) + 2H^*(aq) \rightarrow Zn^{2*}(aq) + H_2(g)$$

With $PH_2 = 1.0$ atm and $[Zn^2+] = 1.0$ M, the cell potential is 0.66 V. The concentration of H* in the cathode compartment is

A) 1.4×10^{-1}

B) 4.9 - 101

C) 4.2 x 10-4

D) 1.0 x 10-12

E) 2.0×10^{-2}

ΔS is negative for the reaction

A) H2O (I) - H2O (g)

B) 25O₂ (g) + O₂ (g) → 25O₃ (g)

C) $NH_4Cl(s) \rightarrow NH_3(g) + HCl(g)$

D) 2C (s) + O₂ (g) → 2CO₂ (g)

E) PbCl₂ (s) \rightarrow Pb²⁺ (aq) + 2Cl⁻ (aq)

5) The standard emf for the cell using the overall cell reaction below is +2.20 V:

2Al(s)+3l2(s) - 2Al3+(aq)+6l-(aq)

The emi generated by the cell when $[A1^{3+}] = 4.5 \times 10^{-3}$ M and $[1^{-}] = 0.15$ M is ______ V. A) 2.23 B) 2.32 C) 2.20 D) 2.10 E) 2.39

6) Which substance is the reducing agent in the reaction below?

Pb + PbO2 + 2H25O4 - 2Pb5O4 + 2H2O

A) H₂SO₄

B) PbO₂

C) Pb

D) H₂O

E) Pb5O₄.

6)

Unit 9 Thermochemistry Test

G Psacharopoulos

Unit 9 Thermochemistry Test:	
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