

Periodic Trends

Period

1. Discuss the importance of Mendeleev's periodic law.
When elements are arranged by increasing number of protons (atomic number), there are repeating patterns of chemical and physical properties.
2. Identify each element as a metal, metalloid, or nonmetal.
 - a) Fluorine **Nonmetal**
 - b) Germanium **Metalloid**
 - c) zinc **Metal**
 - d) phosphorous **Nonmetal**
 - e) lithium **Metal**
3. Give two examples of elements for each category.
 - a) noble gases **He, Ne, Ar, Kr, Xe, Rn, 118**
 - b) halogens **F, Cl, Br, I, At, 117**
 - c) alkali metals **Li, Na, K, Rb, Cs, Fr**
 - d) alkaline earth metals **Be, Mg, Ca, Sr, Ba, Ra**
3. What trend in atomic radius do you see as you go down a group/family on the periodic table? What causes this trend? **Atomic radii INCREASE going down a group due to added energy levels and increases shielding effects.**
4. What trend in atomic radius do you see as you go across a period/row on the periodic table? What causes this trend? **Atomic radii DECREASE going across a period due to added energy levels and increases shielding effects.**
5. Circle the atom in each pair that has the largest atomic radius.
 - a) **Al** B d) Na **Al**
 - b) **S** O e) O **F**
 - c) **Br** Cl f) Mg **Ca**
6. Define ionization energy.
The amount of energy required to remove a valence electron.
7. What trend in ionization energy do you see as you go down a group/family on the periodic table? What causes this trend? **Ionization energy DECREASES as you go down a group, as the atomic radius increases. This is due to greater shielding effects.**

Unit 8 Quiz Periodic Trends Answer Key

Alyn G. McFarland



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