UNIT 5: CHEMICAL KINETICS AND EQUILIBRIUM

Chapter 14: Chemical Kinetics

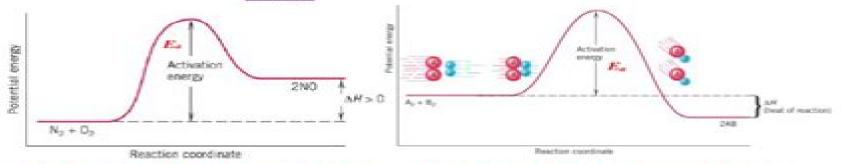
14.4 & 14.6: Activation Energy, Temperature Dependence on Reaction Rates & Catalysis

Reaction Rates - the speed of which the concentration of a reactant or product changes over time.

Collision Model - a model that state for a reaction to occur, molecules must collide with each other.

Factors Affecting the Collision Model

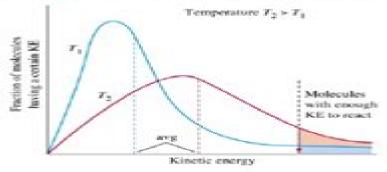
- Activation Energy (E_a) the threshold energy molecules needed to overcome to cause a chemical reaction that was first proposed by Svante Arrhenius
 - E_d is the highest energy (top of the hill E_{proc}) minus the sum of energy of the reactants (ΣΗ_{procussor}) on the potential energy diagram.
 - in general, the Lower the Activation Energy, the Faster is the Rate of Reaction



Activation Energy of an Endothermic Reaction

Activation Energy of an Exothermic Reaction

Temperature (7) - the effective number of collisions increases exponentially with temperature.



Because $T_2 > T_1$, there are more molecules colliding with enough Kinetic Energy causing a reaction to occur.

In general, the Higher the Temperature, the Faster the Rate of Reaction.

Unit 8 Kinetics And Equilibrium Answers

Katrin Zwirglmaier



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