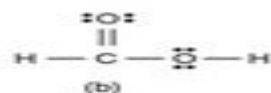
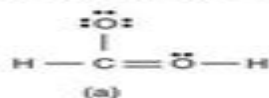


28. Carbon and oxygen can bond either by a single bond (as in $\text{CH}_3\text{-OH}$), a double bond (as in O=C=O), or a triple bond (as in $\text{HC}\equiv\text{CH}$).
- Describe these three types of bonds in terms of σ -bonds and π -bonds.
 - How would you expect the length of the carbon-oxygen bond to vary in the three given examples?
29. Explain the following in terms of the intermolecular forces that exist in them:
- Water is a liquid at room temperature and atmospheric pressure, but hydrogen sulphide is a gas.
 - At room temperature and atmospheric pressure chlorine is a gas, bromine is a liquid and iodine is a solid.
 - Pentan-1-ol boils at 137°C , whereas pentan-3-ol boils at 116°C .
 - The boiling point of sulphur dioxide is 24°C higher than that of chlorine.
30. Of the two possible valence-bond formulas for formic acid (HCO_2H), which one is the more reasonable structure?



31. Draw the correct valence-bond formula for hydrazine (N_2H_4).
32. Draw the resonance hybrid of SO_2 .
33. Based on the VB approach, predict the shape of a ClO_3^- ion.
34. The molecular orbital description of NO is similar to that of O_2 except that the energies of the oxygen atomic orbitals are slightly lower than those of nitrogen atomic orbitals and NO molecule contains different number electrons than O_2 .
- Draw an energy diagram to show how the atomic orbitals of N and O combine to form molecular orbitals of NO.
 - How many unpaired electrons are present in the NO molecule?
 - What is the bond order in NO molecule?
 - Is NO paramagnetic or diamagnetic?
35. List the following in order of increasing O – O bond length : O_2 , O_2^+ , O_2^-
36. Show the direction of the bond moment (dipole moment) in each of the following bonds:



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