

Chemistry Unit 3 Test Review

5.1 Mole and Avogadro's Constant

A **particle** is very small and can be:

- A single atom (Ca)
- A molecule with 2 or more atoms (molecular compound – H₂O)
- An ion with a positive or negative charge (CO₃²⁻)
- A formula unit with two or more ions (ionic compound – NaCl)

The Mole

- Chemists use the mole to represent a specific number of particles
- The number of particles it represents is called the Avogadro constant (6.02×10^{23} particles/mol)
- The unit for mole is mol

Definitions

1. Mole: One mole of any entity is 6.02×10^{23} of the particles (atom, molecule, ion formula unit); the symbol is n; the unit is mol
2. Avogadro's constant: The number of particles in one mole, 6.02×10^{23} ; the symbol is N_A; the unit is particles/mol
3. Particles:
 - Atom or element (Eg. Zn)
 - Molecule or covalent compound (Eg. H₂O)
 - Ions or monatomic/polyatomic ions (Eg. Zn⁺² and CO₃²⁻)
 - Formula unit or ionic compound (Eg. LiCl)

5.2 Molar Mass

Basics: The Mass and The Mole

- The mass of carbon-12 is 12 atomic mass unit (12 u). Every other element has mass relative to one atom of carbon-12
- 1 mole of carbon-12 = 6.02×10^{23} atoms of C-12 = 12g
- The average atomic mass of any element, in atomic mass units, is numerically equivalent to the mass of one mole of atoms (6.02×10^{23} atoms) of that element

Molar Mass: The mass of one mole of its particle – atoms, molecules, ions, or formula units, in grams per mole (g/mol)

6.1 Percentage Composition

Law of Definite Proportions:

- Elements in a chemical compound are always present in the same proportions by mass
- *Law of constant composition:* all molecules or formula units of a compound will have the same chemical composition

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Table of Contents Unit 3 Chemistry Test Answers Science Explorer

1. Understanding the eBook Unit 3 Chemistry Test Answers Science Explorer
 - The Rise of Digital Reading Unit 3 Chemistry Test Answers Science Explorer
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 3 Chemistry Test Answers Science Explorer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 3 Chemistry Test Answers Science Explorer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 3 Chemistry Test Answers Science Explorer
 - Personalized Recommendations
 - Unit 3 Chemistry Test Answers Science Explorer User Reviews and Ratings
 - Unit 3 Chemistry Test Answers Science Explorer and Bestseller Lists
5. Accessing Unit 3 Chemistry Test Answers Science Explorer Free and Paid eBooks

- Unit 3 Chemistry Test Answers Science Explorer Public Domain eBooks
- Unit 3 Chemistry Test Answers Science Explorer eBook Subscription Services
- Unit 3 Chemistry Test Answers Science Explorer Budget-Friendly Options
- 6. Navigating Unit 3 Chemistry Test Answers Science Explorer eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 3 Chemistry Test Answers Science Explorer Compatibility with Devices
 - Unit 3 Chemistry Test Answers Science Explorer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 3 Chemistry Test Answers Science Explorer
 - Highlighting and Note-Taking Unit 3 Chemistry Test Answers Science Explorer
 - Interactive Elements Unit 3 Chemistry Test Answers Science Explorer
- 8. Staying Engaged with Unit 3 Chemistry Test Answers Science Explorer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 3 Chemistry Test Answers Science Explorer
- 9. Balancing eBooks and Physical Books Unit 3 Chemistry Test Answers Science Explorer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 3 Chemistry Test Answers Science Explorer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 3 Chemistry Test Answers Science Explorer
 - Setting Reading Goals Unit 3 Chemistry Test Answers Science Explorer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 3 Chemistry Test Answers Science Explorer
 - Fact-Checking eBook Content of Unit 3 Chemistry Test Answers Science Explorer
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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