

# CHEMISTRY

## UNIT 3: RATES, EQUILIBRIUM & ORGANIC CHEM

### Rates & equilibrium (C8):

#### Rate of reaction

- The rate of reactions tells you how fast reactants turn into products
- Reactions happen at all sorts of different rates, some are fast like combustion, some are slow like rusting
- You can find the rate of reaction as: amount of reactant/product used

Time

- Quantity of reactant/product can be measured in grams, moles or, if volume,  $\text{cm}^3$
- To find the rate of reaction graphically, measure the rate of reaction, draw tangents to curve and measure the rate of reaction from the tangent

- To monitor mass of a reaction mixture, place the reaction flask on a balance. In these reactions, a gas is given off, so record the decrease in mass in time intervals. Plot a graph of time vs mass

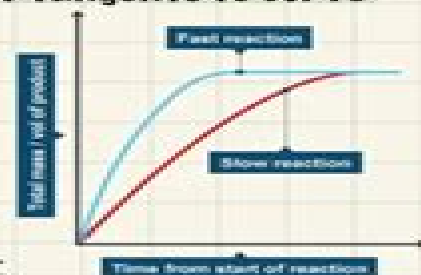
- To monitor volume of gas. Connect a syringe to the reaction flask and measure the volume formed in time intervals. Plot a graph of volume vs time.

- To monitor cross with sodium thiosulphate:

- 1) use a measuring cylinder to add 10 ml of sodium thiosulfate solution to a conical flask, then add 40 ml of water
- 2) measure and record the gas
- 3) place the conical flask on a piece of paper with a black cross drawn on it
- 4) use another cylinder to add 10 ml of hydrochloric acid to the flask, and start a timer
- 5) when the cross is no longer visible record the time taken, and then clean the apparatus
- 6) repeat steps 1-5 changing the concentration of sodium thiosulphate and measure the mean time

control variables: volume and concentration of water and hydrochloric acid

Volume of gas produced ( $\text{cm}^3$ )



# Unit 16 Reaction Rates Notes

**RM Cervero**



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