

Unit 10 Stoichiometry Test 1: Short Answers

Answer all questions including units and balanced equations where appropriate.

1. How many moles of aluminium are produced when completely with 2.5 mol of FeCl_3 ?

$$2\text{Al} + 3\text{FeCl}_3 \rightarrow 3\text{Fe} + \text{Al}_2\text{Cl}_6$$
2. How many grams of FeCl_3 are produced when 2.5 moles of aluminium with oxygen to form Al_2O_3 ?

$$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$$
3. The reaction below shows that decomposition of Hg_2Cl_2 occurs. How many grams of oxygen are produced when 20.7g Hg_2Cl_2 is heated?

$$\text{Hg}_2\text{Cl}_2 \rightarrow 2\text{Hg} + \text{Cl}_2$$
4. An element that has been suggested to produce 2.0g of HgCl_2 actually produces 2.2g of HgCl_2 . What is the percent yield?
5. How many moles of oxygen are needed to produce 20.0g of Fe_2O_3 when 20.0g of iron reacts with O_2 ?

$$4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$$

6. Percent Yield is

7. Amount of 20.0g produced is

8. How actually when 20.0g of Fe_2O_3 what is the percent yield?

Unit 10 Stoichiometry Test Review

Alen Hadzovic

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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