

Equilibrium States

1. A thermodynamic system is initially in equilibrium, and there are four thermodynamic changes, separately, greater which observations/measurements can still find the system in equilibrium or, more easily, spontaneously, forward, backward, neither, nor change.

$$H_{\text{sys}} = 4.5 \text{ kJ}, S_{\text{sys}} = 1.0 \text{ J/K}$$

a. The pressure of the system is decreased

forward

b. The pressure of the system is decreased

forward

c. The pressure of the system is decreased

forward

2. A thermodynamic system is initially in equilibrium, and there are a total of eight, positive which observation the measurements will be found in equilibrium or more easily, spontaneously, forward, backward, neither, nor change.

$$\text{for } H_{\text{sys}} = 1.0 \text{ kJ}, S_{\text{sys}} = 1.0 \text{ J/K}$$

The pressure and the volume of the system are decreased, and the temperature is increased.

3. With the observations:

$$H_{\text{sys}} = 1.0 \text{ kJ}, S_{\text{sys}} = 1.0 \text{ J/K}$$

a. All the factors in these states (the equilibrium representation H_{sys})

$$\frac{H_{\text{sys}}}{S_{\text{sys}}} = \frac{1.0 \text{ kJ}}{1.0 \text{ J/K}}$$

b. All the factors in these states are the same, $H_{\text{sys}} = 1.0 \text{ kJ}$, $S_{\text{sys}} = 1.0 \text{ J/K}$, and $T_{\text{sys}} = 1.0 \text{ K}$.

$$\frac{H_{\text{sys}}}{S_{\text{sys}}} = \frac{1.0 \text{ kJ}}{1.0 \text{ J/K}}$$

$$\frac{H_{\text{sys}}}{S_{\text{sys}}} = \frac{1.0 \text{ kJ}}{1.0 \text{ J/K}} = 1.0 \text{ kJ/K}$$

Systems At Equilibrium Concept Review Answers

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