

UNIT 1 KEY TERMS

TOPIC 1.1 - STRUCTURE OF WATER & HYDROGEN BONDING

Acid:

A substance that increases the concentration of hydrogen ions (H^+) in a solution.

Adhesion:

The attraction between water molecules and other substances.

Base:

A substance that provides hydroxide ions (OH^-), reducing the concentration of hydrogen ions in a solution.

Buffer:

A substance that can resist changes in pH by donating excess hydrogen ions (H^+) or hydroxide ions (OH^-).

Capillary Action:

The ability of a liquid to flow in reverse against the force of gravity without the assistance of, or against, external forces like gravity.

Cohesion:

The attraction between water molecules due to hydrogen bonding. It allows water molecules to stick together.

Dissociation:

Occurs when atoms or groups of atoms break off from molecules and form ions.

Electronegativity:

A measure of an atom's ability to attract and hold electrons.

Heat of Vaporization of Water:

The amount of energy required to change one gram of liquid water into vapor at constant temperature and pressure.

Hydrogen Bond:

A weak electrostatic attraction between a partially positively charged hydrogen atom and a partially negatively charged atom (usually O or N) on the another molecule.

Hydrophilic:

Substances that have an affinity for water and can interact or dissolve in it. They are typically polar or charged.

Hydrophobic:

Substances that repel water and do not dissolve well in it. They are often non-polar or lack charged regions.

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