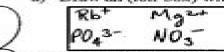
- An aqueous solution of rubidium phosphate (Rb<sub>2</sub>PO<sub>4</sub>) is mixed with an aqueous solution of magnesium nitrate and a precipitate forms.
  - a) Draw an (ion-box) with the four ions involved with this reaction. (1 mark)



b) Write a balanced formula equation for this reaction. Include all subscripts and coefficients. (2 marks)

2 Rb 3 PO+(48) +3Mg(NO3)2(49) -> Mg (PO4)2(3) + 6 Rb NO3(48)

 Write a balanced, complete (total) tonic equation for this reaction. Include all subscripts, coefficients and ion charges where necessary. (2 marks)

6 Rb (og ,+ 2 PO + 3 Mg 2 tog + 6 NO 5 tog, - Mg (PO+) 2(5) + 6 Rb (og) + 6 NO 3 (og)

 Write a net ionic equation for this reaction. Include all subscripts, coefficients and ion charges where necessary. (2 marks)

7

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**Staley Wilbraham (Matta, Waterman)** 

**Tesccc Chemistry Unit 6 Lesson Answers:** 

**Pearson Chemistry Unit 6** Staley Wilbraham (Matta, Waterman),2022

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