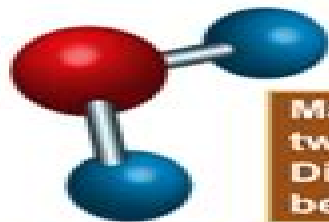


Name: _____



Substances vs. Mixtures

Matter can either be a substance or a mixture. Oftentimes, these two categories of matter are confused with one another. Differentiate the two categories of matter through the table below.

Compare and contrast substances and mixtures through the characteristics listed below. Put a check (✓) mark under the category through which it applies. Then, answer the question that follows. You may do additional research to accomplish this activity.

Characteristics	Substances	Mixtures
Homogeneous		
Heterogeneous		
Fixed Composition		
Fixed Boiling Point		
Fixed Melting Point		
Chemically-combined		
Held by chemical bonds		
Made up of 1 or more atoms		
Made up of 2 or more compounds		

HOW ARE SUBSTANCES DIFFERENT FROM MIXTURES?



Substances Mixtures And Solubility Answer Sheet

Stacey Faulkner



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Qualitative Organic Microanalysis Frank Schneider, 2012-12-06 The most important task of the analytical chemist aside from the acquisition of experimental data is the coordination and interpretation of such data in terms of the qualitative and quantitative composition of the test substance As in the old tale of the blind men and the elephant a single observation or test not considered in conjunction with others may lead to entirely erroneous conclusions On the other hand mere increase in the number of such tests without regard to their need or to their relationship to each other also may not suffice for drawing the correct inferences from the experimental evidence The deductive reasoning which is usually associated with the analytical chemist finds its greatest opportunity for application in the problems of cognition and recognition of carbon compounds Since a rigid scheme of procedure tends to produce a corresponding rigidity of thought the intent of the present book is to outline approaches which will minimize the chances of misinterpretation without restricting the analytical chemist in his choice of tests The selection of subsequent tests or reactions should be governed by the results of the preceding ones rather than by an arbitrary list The relationships of the various approaches and the information derived from each to the composition and constitution of the test substance are brought out in the discussion with the hope that they may serve as guide lines for such selection

Thermodynamics of Solutions Eli Ruckenstein, Ivan L. Shulgin, 2009-06-17 This book consists of a number of papers regarding the thermodynamics and structure of multicomponent systems that we have published during the last decade Even though they involve different topics and different systems they have something in common which can be considered as the signature of the present book First these papers are concerned with difficult or very nonideal systems i e systems with very strong interactions e g hydrogen bonding between components or systems with large differences in the partial molar volumes of the components e g the aqueous solutions of proteins or systems that are far from normal conditions e g critical or near critical mixtures Second the conventional thermodynamic methods are not sufficient for the accurate treatment of these mixtures Last but not least these systems are of interest for the pharmaceutical biomedical and related industries In order to meet the thermodynamic challenges involved in these complex mixtures we employed a variety of traditional methods but also new methods such as the fluctuation theory of Kirkwood and Buff and ab initio quantum mechanical techniques The Kirkwood Buff KB theory is a rigorous formalism which is free of any of the approximations usually used in the thermodynamic treatment of multicomponent systems This theory appears to be very fruitful when applied to the above mentioned difficult systems

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