

Earth Science Reference Tables

PHYSICAL CONSTANTS

Radioactive Decay Data

RADIOACTIVE ISOTOPE	DISINTEGRATION	HALF-LIFE (years)
Carbon-14	$C^{14} \rightarrow N^{14}$	5.7×10^3
Potassium-40	$K^{40} \rightarrow Ar^{40}$ $K^{40} \rightarrow Ca^{40}$	1.3×10^9
Uranium-238	$U^{238} \rightarrow Pb^{206}$	4.5×10^9
Rubidium-87	$Rb^{87} \rightarrow Sr^{87}$	4.9×10^{10}

Specific Heats of Common Materials

MATERIAL	S-SPECIFIC HEAT (calories/gram · °C)
Water { solid	0.5
liquid	1.0
gas	0.5
Dry air	0.24
Basalt	0.20
Granite	0.19
Iron	0.11
Copper	0.09
Lead	0.03

Properties of Water

Energy gained during melting	80 calories/gram
Energy released during freezing	80 calories/gram
Energy gained during vaporization	540 calories/gram
Energy released during condensation	540 calories/gram
Density at 3.98°C	1.00 gram/milliliter

EQUATIONS

Percent deviation
from accepted value

$$\text{deviation (\%)} = \frac{\text{difference from accepted value}}{\text{accepted value}} \times 100$$

Eccentricity of an ellipse

$$\text{eccentricity} = \frac{\text{distance between foci}}{\text{length of major axis}}$$

Gradient

$$\text{gradient} = \frac{\text{change in field value}}{\text{distance}}$$

Rate of change

$$\text{rate of change} = \frac{\text{change in field value}}{\text{time}}$$

Density of a substance

$$\text{density} = \frac{\text{mass}}{\text{volume}}$$

2001 EDITION

This edition of the Earth Science Reference Tables should be used in the classroom beginning in the 2000-2001 school year. The first examination for which these tables will be used is the January 2001 Regents Examination in Earth Science.


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Use Earth Science Reference Table

SJ Ball



Use Earth Science Reference Table:

Decoding **Use Earth Science Reference Table**: Revealing the Captivating Potential of Verbal Expression

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