Name: __KEY______ Section: _____

Chapter 7 Practice Worksheet: Covalent Bonds and Molecular Structure

1) How are ionic bonds and covalent bonds different?

Ionic bonds result from the transfer of electrons from one atom to another, Covalent bonds result from two atoms sharing electrons.

Describe the relationship between the length of a bond and the strength of that bond.

Strength of a bond increases as the bond gets shorter (inverse relationship)

Identify the type(s) of bond(s) found in the following molecules:

a. CCl₄ covalent

b. Li₂O ionic

c. NF₃ covalent_____

d. CaSO₄ jonic and covalent

e. SO₂ covalent

f. Mg(OH)₂ ionic and covalent

4) Define electronegativity.

Electronegativity is the ability for an atom in a molecule to attract electrons to itself.

 Use electronegativity values to place the following elements in increasing order: F, N, H, P, Si, C, O

 $Si < C \approx P = H < N < O < F$

6) Determine if the bond between atoms in each example below is nonpolar covalent, polar covalent, or ionic.

a. H₂ npc

e. NF __pc___

b. PCI __pe____

f. MgO ionic

e. F₂ npe____

g. CH __npe__

d NaBr ionic

h. HCl pc

<u>Using Science Skills Relating Concepts Chapter 7</u> <u>Answers</u>

Lisa A. Baglione

Using Science Skills Relating Concepts Chapter 7 Answers:

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