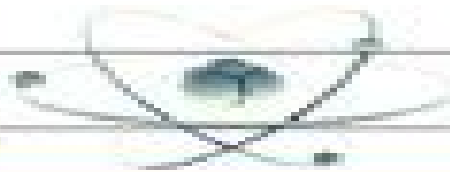


Matter and Energy



Classification of Matter (N.5.A.1)

Mass: Measure of the amount of matter in an object, typically in grams or kilograms.

Magnetism: Ability of a substance to attract or be attracted to other magnetic materials.

Physical States:

- **Solid:** Fixed shape and volume.
- **Liquid:** Fixed volume but takes the shape of its container.
- **Gas:** Neither fixed shape nor volume, expands to fill its container.

Relative Density: Determines whether a substance will sink (denser than the fluid) or float (less dense than the fluid).

Solubility in Water: Ability of a substance to dissolve in water.

Thermal and Electrical Conductivity:

Conductors: Materials that allow heat or electricity to flow through them easily.

Insulators: Materials that do not allow heat or electricity to flow through them easily.

Boiling and Freezing Points of Water

Boiling Point: The temperature at which water turns into steam (100°C at sea level).

Freezing/Melting Point: The temperature at which water becomes ice (0°C).

Physical Properties of Mixtures

Examples of Mixtures:

- **Iron Filings and Sand:** A mixture where both components retain their individual properties, e.g., magnetism of iron filings.

Changes in States of Matter

Evaporation:

Causes liquid to melt and liquids to evaporate depending on the temperature reached.

Example: Ice melting and water evaporating into steam.

Condensing:

Causes gases to condense and liquids to freeze if cooled sufficiently.

Example: Steam condensing into water, water freezing into ice.

Changes in Physical Properties of Solutions
Dissolving: Salt in Water
 Salt (sodium chloride) when dissolved in water becomes aqueous, making the solution conductive.
Adding Lemon Juice to Water: Changes the taste and pH of the solution, demonstrating a change in the physical properties.

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