

Derive the following equations:

1. Show the following thermodynamic equations:

$$d(\Delta G_{\text{mix}}) = Td(\Delta S_{\text{mix}}) - d(\Delta H_{\text{mix}}) + d(\Delta PV_{\text{mix}}) = -RTd\ln Q$$

(a) is this equation a thermodynamic or mathematical relationship?

$$d(\Delta G_{\text{mix}}) = d(\Delta H_{\text{mix}}) - Td(\Delta S_{\text{mix}}) + d(\Delta PV_{\text{mix}})$$

(b) what equation the mixed partials of thermodynamic quantities derived?

$$d(\Delta G_{\text{mix}}) = \frac{\partial(\Delta G_{\text{mix}})}{\partial T} dT = -\Delta S_{\text{mix}} dT$$

(c) what equation the mixed partials of thermodynamic quantities derived?

$$\Delta G_{\text{mix}} = \frac{\partial(\Delta G_{\text{mix}})}{\partial P} = \Delta V_{\text{mix}} = \Delta V_{\text{mix}} dP$$

(d) what equation the mixed partials of thermodynamic quantities derived?

$$\Delta G_{\text{mix}} = \frac{\partial(\Delta G_{\text{mix}})}{\partial \ln Q} = \frac{\partial(\Delta G_{\text{mix}})}{\partial \ln Q} = -RT \ln Q$$

2. Thermodynamic equations: $d(\Delta G_{\text{mix}}) = Td(\Delta S_{\text{mix}}) - d(\Delta H_{\text{mix}}) + d(\Delta PV_{\text{mix}}) = -RTd\ln Q$

is this thermodynamic equation which must include the appropriate sign with the appropriate sign for the thermodynamic quantities equation

$$d(\Delta G_{\text{mix}}) = Td(\Delta S_{\text{mix}}) - d(\Delta H_{\text{mix}}) + d(\Delta PV_{\text{mix}}) = -RTd\ln Q$$

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$$d(\Delta G_{\text{mix}}) = Td(\Delta S_{\text{mix}}) - d(\Delta H_{\text{mix}}) + d(\Delta PV_{\text{mix}}) = -RTd\ln Q$$

$$d(\Delta G_{\text{mix}}) = \frac{\partial(\Delta G_{\text{mix}})}{\partial T} dT = -\Delta S_{\text{mix}} dT$$

is this thermodynamic equation which must include the appropriate sign with the appropriate sign for the thermodynamic quantities equation

$$d(\Delta G_{\text{mix}}) = \frac{\partial(\Delta G_{\text{mix}})}{\partial P} dP = \Delta V_{\text{mix}} dP$$

is this thermodynamic equation which must include the appropriate sign with the appropriate sign for the thermodynamic quantities equation

$$d(\Delta G_{\text{mix}}) = \frac{\partial(\Delta G_{\text{mix}})}{\partial \ln Q} d\ln Q = -RT d\ln Q$$

3. Thermodynamic equations: $d(\Delta G_{\text{mix}}) = Td(\Delta S_{\text{mix}}) - d(\Delta H_{\text{mix}}) + d(\Delta PV_{\text{mix}}) = -RTd\ln Q$

a. Show a thermodynamic equation for the derivation.

$$d(\Delta G_{\text{mix}}) = Td(\Delta S_{\text{mix}}) - d(\Delta H_{\text{mix}}) + d(\Delta PV_{\text{mix}}) = -RTd\ln Q$$

is this thermodynamic equation which must include the appropriate sign with the appropriate sign for the thermodynamic quantities equation

$$d(\Delta G_{\text{mix}}) = \frac{\partial(\Delta G_{\text{mix}})}{\partial T} dT = -\Delta S_{\text{mix}} dT$$

is this thermodynamic equation which must include the appropriate sign with the appropriate sign for the thermodynamic quantities equation

$$d(\Delta G_{\text{mix}}) = \frac{\partial(\Delta G_{\text{mix}})}{\partial P} dP = \Delta V_{\text{mix}} dP$$

Study Guide Thermochemical Equations Answers

Lingjun Ying



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