Stoichiometry, Limiting Reactants, Percent Yield, and Excess Reactant Remaining Worksheet

 Aluminum reacts with oxygen to produce aluminum oxide which can be used as an adsorbent, desiccant or catalyst for organic reactions.

> 1 mole of aluminum (Al) is 26.982 g 1 mole of molecular oxygen (O₂) is 31.998 g 1 mole of aluminum oxide is 101.961 g

a. What is the formula of aluminum oxide? (First hint: use the criss-cross rule with aluminum ion and oxide ion) (Second hint: check your answer with the molar mass of aluminum oxide)

ALO

b. Write the balanced chemical equation for the reaction of aluminum (Al) with molecular oxygen (O₂) to make aluminum oxide.

c. How many grams of aluminum oxide could form if 82.49 g Al reacts completely?

82.49 gA1 x
$$\frac{1 \, mol \, Al}{26.982 \, g \, Al}$$
 x $\frac{2 \, mol \, Al_2 \, O_3}{4 \, mol \, Al}$ x $\frac{101.961 \, g \, Al_2 \, O_3}{1 \, mol \, Al_2 \, O_3} = 155.9 \, g \, Al \, O_3$

d. How many grams of aluminum oxide could form if 117.65 g O2 reacts completely?

117.65 g
$$O_2$$
 x $\frac{1 \, mol \, O_2}{31.998 \, g \, O_3}$ x $\frac{2 \, mol \, Al_2 \, O_3}{3 \, mol \, O_2}$ x $\frac{101.961 \, g \, Al_2 \, O_3}{1 \, mol \, Al_2 \, O_3} = 249.9 \, g \, Al_2 \, O_3$

Study Guide Stoichiometry Limiting Reactant Percent Yield

Richard Bailey

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Table of Contents Study Guide Stoichiometry Limiting Reactant Percent Yield

- 1. Understanding the eBook Study Guide Stoichiometry Limiting Reactant Percent Yield
 - The Rise of Digital Reading Study Guide Stoichiometry Limiting Reactant Percent Yield
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Study Guide Stoichiometry Limiting Reactant Percent Yield
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Stoichiometry Limiting Reactant Percent Yield
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Study Guide Stoichiometry Limiting Reactant Percent Yield
 - Personalized Recommendations
 - Study Guide Stoichiometry Limiting Reactant Percent Yield User Reviews and Ratings
 - Study Guide Stoichiometry Limiting Reactant Percent Yield and Bestseller Lists
- 5. Accessing Study Guide Stoichiometry Limiting Reactant Percent Yield Free and Paid eBooks
 - Study Guide Stoichiometry Limiting Reactant Percent Yield Public Domain eBooks
 - Study Guide Stoichiometry Limiting Reactant Percent Yield eBook Subscription Services
 - Study Guide Stoichiometry Limiting Reactant Percent Yield Budget-Friendly Options

- 6. Navigating Study Guide Stoichiometry Limiting Reactant Percent Yield eBook Formats
 - o ePub, PDF, MOBI, and More
 - Study Guide Stoichiometry Limiting Reactant Percent Yield Compatibility with Devices
 - Study Guide Stoichiometry Limiting Reactant Percent Yield Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Study Guide Stoichiometry Limiting Reactant Percent Yield
 - Highlighting and Note-Taking Study Guide Stoichiometry Limiting Reactant Percent Yield
 - Interactive Elements Study Guide Stoichiometry Limiting Reactant Percent Yield
- 8. Staying Engaged with Study Guide Stoichiometry Limiting Reactant Percent Yield
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide Stoichiometry Limiting Reactant Percent Yield
- 9. Balancing eBooks and Physical Books Study Guide Stoichiometry Limiting Reactant Percent Yield
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Study Guide Stoichiometry Limiting Reactant Percent Yield
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide Stoichiometry Limiting Reactant Percent Yield
 - Setting Reading Goals Study Guide Stoichiometry Limiting Reactant Percent Yield
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide Stoichiometry Limiting Reactant Percent Yield
 - Fact-Checking eBook Content of Study Guide Stoichiometry Limiting Reactant Percent Yield
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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