



Chemistry

Paper Review and Study Guide

Periodic Table

1. Explain how the number of protons in an atom's nucleus...

2. Describe the trend in electronegativity the periodic table and explain the reason...

3. How does the periodic table help...

Periodic Table and Law

Development of the Modern Periodic Table

Early Periodic Tables:

- Dmitri Mendeleev in 1869, Dmitri Mendeleev published the first periodic table. He arranged elements by increasing atomic mass and noticed that certain properties periodically repeated. Mendeleev's table left gaps for elements that had not yet been discovered, predicting their properties accurately.
- Luther Meyer independently, Luther Meyer published a similar periodic table around the same time as Mendeleev.

Modern Periodic Table:

- Moseley's Discovery: In 1913, Henry Moseley discovered the atomic number of each element, leading to the modern definition of the periodic law. The modern periodic table is arranged by increasing atomic number rather than atomic mass.

The periodic table is a powerful tool for understanding the behavior of elements. It's organized based on atomic numbers, with elements arranged in rows called periods and columns called groups. Elements within the same group share similar properties, while elements in the same period have the same number of electron shells. The elements are separated into metals, nonmetals, and metalloids.

Periodic Law:

The periodic law states that the properties of the elements are a function of their atomic number. This means that when elements are arranged by increasing atomic number, elements with similar properties appear at regular intervals.

Periods and Groups:

Periods are horizontal rows of elements. There are 7 periods in the modern periodic table. Each period corresponds to a principal energy level of an atom.

Groups are vertical columns on the periodic table. Elements in the same group have similar chemical properties.

Periodic Table and Law

4-page Review +

39 Question Study Guide

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O. Ray Whittington



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