

# 20 Oxidation-Reduction Reactions

## Big idea REACTIONS

### 20.1 The Meaning of Oxidation and Reduction

For students using the Foundation edition, assign problems 2–6, 8, 10–16, 18, 19.

**Essential Understanding** Oxidation and reduction are opposite chemical processes during which electrons are shifted between reactants.

#### Reading Strategy

**Compare and Contrast** Organizing information in a table helps you compare and contrast several topics at one time. Consider comparing and contrasting oxidation and reduction and how they relate to oxygen and electrons. As you read, ask yourself, “How are they similar? How are they different?”

As you read Lesson 20.1, use the compare and contrast table below. Complete the table, comparing how oxidation and reduction relate to electrons and oxygen.

| Oxidation and Reduction |                                                                                             |                                                                                                                              |
|-------------------------|---------------------------------------------------------------------------------------------|------------------------------------------------------------------------------------------------------------------------------|
|                         | Compare                                                                                     | Contrast                                                                                                                     |
| Oxygen                  | <i>If oxygen is involved in the process, a change occurs in the number of oxygen atoms.</i> | <i>During oxidation, a substance gains oxygen. During reduction, a substance loses oxygen.</i>                               |
| Electrons               | <i>A change occurs in the number of electrons.</i>                                          | <i>During oxidation, a substance loses one or more electrons. During reduction, a substance gains one or more electrons.</i> |

**EXTENSION** Use the information about electrons in the chart to explain why oxidation and reduction always occur together.

*When electrons are gained by one substance, the electrons must come from another substance, which loses the electrons.*

# Study Guide Redox Reaction Answers

**Seager/Slabaugh**



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**Chemical Principles Study Guide/Solutions Manual** John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications

**Chemical Principles Student's Study Guide & Solutions Manual** John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text

**Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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