

Gas Stoichiometry Chemistry 110

1] Given the equation: $2 \text{NH}_3(\text{g}) + 3 \text{Cl}_2(\text{g}) \rightarrow \text{N}_2(\text{g}) + 6 \text{HCl}(\text{g})$

a. How many milliliters of nitrogen can be made from 13 L of chlorine and 10.0 L of ammonia gas at STP?

$$10.0 \text{ L NH}_3 \times \frac{1 \text{ L N}_2}{2 \text{ L NH}_3} = 5.00 \text{ L N}_2$$

$$13 \text{ L Cl}_2 \times \frac{1 \text{ L N}_2}{3 \text{ L Cl}_2} = 4.3 \text{ L N}_2$$

Answer 4.3 x 10³ mL N₂

b. How many grams of chlorine must react to produce 16 L of nitrogen gas at 1.2 atm and 23°C?

$$n_{\text{N}_2} = \frac{(1.2 \text{ atm}) \times (16 \text{ L})}{(0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}) \times (273 + 23) \text{ K}} = 0.790 \text{ mol N}_2$$

$$0.790 \text{ mol N}_2 \times \frac{3 \text{ mol Cl}_2}{1 \text{ mol N}_2} \times \frac{70.1 \text{ g Cl}_2}{1 \text{ mol}} = 1.66 \times 10^2 \text{ g Cl}_2$$

Answer 1.66 x 10² g Cl₂

c. How many liters of ammonia gas at 244 torr and 35°C must be used to produce 2.3 kg of HCl gas?

$$2.3 \text{ kg HCl} \times \frac{10^3 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mol HCl}}{36.5 \text{ g}} \times \frac{2 \text{ mol NH}_3}{6 \text{ mol HCl}} = 21 \text{ mol NH}_3$$

$$V_{\text{NH}_3} = \frac{(21 \text{ mol NH}_3) \times (0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}) \times (273 + 35) \text{ K}}{\frac{244 \text{ torr}}{760 \text{ torr/1 atm}}} = 1.7 \times 10^3 \text{ L NH}_3$$

Answer 1.7 x 10³ L NH₃

2] How many liters of ammonia, measured at STP, must be used to produce of 2.65 grams of calcium hydride 6



$$2.65 \text{ g CaH}_2 \times \frac{1 \text{ mol CaH}_2}{42.1 \text{ g}} \times \frac{2 \text{ mol NH}_3}{3 \text{ mol CaH}_2} \times \frac{22.4 \text{ L}}{1 \text{ mol}} = 0.940 \text{ L NH}_3$$

Answer 0.940 L NH₃

3] What volume of oxygen, measured at 35°C and 752 mm Hg, is required to "burn" 3.26 grams of calcium.



$$3.26 \text{ g Ca} \times \frac{1 \text{ mol Ca}}{40.1 \text{ g}} \times \frac{1 \text{ mol O}_2}{2 \text{ mol Ca}} = 0.0406 \text{ mol O}_2$$

$$V_{\text{O}_2} = \frac{(0.0406 \text{ mol O}_2) \times (0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}) \times (273 + 35) \text{ K}}{\frac{752 \text{ mm}}{760 \text{ mm/1 atm}}} = 1.04 \text{ L O}_2$$

Answer 1.04 L O₂

4] How many mLs of 0.0246 M AgNO₃ required to precipitate as silver chromate all the chromate ion in a solution containing 2.10 g sodium chromate

Sodium chromate + Silver nitrate → Silver chromate (solid) + Sodium nitrate



$$2.10 \text{ g Na}_2\text{CrO}_4 \times \frac{1 \text{ mol Na}_2\text{CrO}_4}{161.8 \text{ g}} \times \frac{2 \text{ mol AgNO}_3}{1 \text{ mol Na}_2\text{CrO}_4} \times \frac{1000 \text{ mL AgNO}_3}{0.0246 \text{ mol}} = 1.05 \times 10^3 \text{ mL}$$

Answer 1.05 x 10³ mL

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A Loxley



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