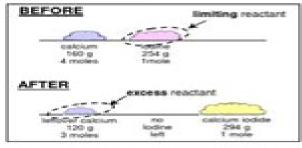
Name _____

When an automobile production plant runs out of tires no more cars can be produced even if there are still plenty of engines, bodies, seats, and other parts. When one of the reactants in a chemical reaction runs out the reaction stops as well. The substance that limits production in a chemical reaction is known as the **limiting** reactant. The reactant that is left over when the reaction stops is called the excess reactant.

1 Ca + 1 I₂ → 1 CaI₂

Consider the chemical reaction between calcium and iodine. When 160 g (4 moles) of calcium reacts with 254 g (1 mole) of iodine the reaction makes 294 g (1 mole) of calcium iodide. Although there is enough calcium to make more calcium iodide, the iodine runs out first. Since all of the iodine gets used up it is called the limiting reactant. According to the balanced equation, if one mole of iodine reacts, one mole of calcium will react. This means that there are still 3 moles of calcium left. Because calcium is left over it is called the excess reactant.



Example

What mass of iron (II) sulfide will be produced if 9.68 g of iron reacts with 6.28 g of sulfur? __Fe + __S → __FeS

- balance the equation <u>1</u> Fe + <u>1</u> S → <u>1</u> FeS
- perform a calculation for each reactant: $p: \frac{9.68 \text{ g/Gc}}{1} \times \frac{1 \text{ m/h/fc}}{55.85 \text{ g/Gc}} \times \frac{1 \text{ m/h/fc}}{1 \text{ m/h/fc}} \times \frac{87.91 \text{ g/Fc}}{1 \text{ m/h/fc}} = 15.24 \text{ g/Fc}$
 - S: 6.28 g/s 1 miles 1 mol Fes | 87.91 g FeS | 17.22 g FeS | 1 mol Fes | 1 mol Fes |
- whichever reactant makes less product is
 the limiting reactant. There is enough sulfur to make 17.22 g
 feS, but only enough iron to make 15.24 g of FeS.

Find the amount of product formed in each of the following reactions.

- What mass of H₂O will be produced if 9.5 g of H₂ reacts with 1.2 g of O₂?
 ___ H₂ + ___ O₂ → ___ H₂O
- If 1.85 g of Mg(OH)₂ reacts with 3.71 g of HCl, how much MgCl₂ is produced? What is the limiting reactant? __Mg(OH)₂ + __HCl → __MgCl₂ + __H₂O
- What mass of AgCl is produced when 53.42 g of AgNO₃ reacts with 14.19 g of NaCl? __AgNO₃ + __NaCl → __AgCl + __NaNO₃
- If 14.7 g of calcium is placed in 11.5 g of water, what mass of hydrogen gas is produced? __Ca + __H₂O → __Ca(OH)₂ + __H₂
- What mass of potassium hydroxide is formed when 8.2 g of potassium oxide is added to 1.3 g of water? ___ K₂O + __ H₂O → __ KOH

- What mass of aluminum chloride could be made from 8.1 g of aluminum and 4.2 L of chlorine at STP? __ Al + __ Cl₂ → __ AlCl₃
- If 5.26 L of nitrogen monoxide and 7.64 L of oxygen are combined, what mass of nitrogen dioxide is formed? Assume conditions are STP.
 NO + _ O₂ → _ NO₂
- If 18.1 g of silicon tetrachloride reacts with 8.4 L of hydrogen at STP, what mass of silicon is formed? __SiCl₄ + __H₂ → __Si + __HCl
- What mass of Cu(OH)₂ is produced in the reaction of 7.6 g Cu(NO₃)₂ with 6.2 g of KOH? _ Cu(NO₃)₂ + _ KOH → _ Cu(OH)₂ + _ KNO₃
- What mass of carbon dioxide is formed when 64 kg of ethylene is burned in 142 kg of oxygen?
 _ C₂H₄ + _ O₂ → _ CO₂ + _ H₂O

Ws 12 3 Limiting Reactants

Chris Sunley

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Cambridge IGCSETM Chemistry Teacher's Guide (Collins Cambridge IGCSETM) Chris Sunley, 2022-02-03 Prepare students with complete coverage of the revised Cambridge IGCSETM Chemistry syllabus 0620 0971 for examination from 2023 Collins Cambridge IGCSE Chemistry Teacher's Guide is full of lesson ideas practical instructions technician s notes planning support and more Chemical Engineering Progress, 1959 Particulate Carbon Donald Siegla.2013-11-11 The goal of the symposium Particulate Carbon Formation During Combustion held at the General Motors Research Laboratories on October 15 and 16 1980 was to discuss fundamental aspects of soot formation and oxidation in combustion systems and to stimulate new research by extensive interactions among the participants This book contains lhe papers and discussions of that symposium the 26th in an annual series covering many different disciplines which are timely and of interest to both General Motors and the technical community at large The subject of this symposium has considerable relevance for man in his effort to control and preserve his environment Emission of particulate carbon into the atmos phere from combustion sources is of concern to scientists and laymen alike The hope of reducing this emission clearly requires an understanding of its formation during the combustion process itself an area of considerable long term research interest It is our hope that this symposium has served to summarize what is known so that what remains to be learned can be pursued Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 with greater vigor Technology, Volume 3 Hisashi Yamamoto, Takashi Kato, 2019-08-05 Edited by foremost leaders in chemical research together with a number of distinguished international authors this third volume summarizes the most important and promising recent developments in material science in one book Interdisciplinary and application oriented this ready reference focuses on innovative methods covering new developments in photofunctional materials polymer chemistry surface science and more Of great interest to chemists as well as material scientists alike Holt Chemistry Ralph Thomas Myers, 2004 Holt **Chemistry** Holt Rinehart & Winston, 2003-01-24 The Educational Value of Chemical Demonstrations in the College Prep Chemistry Classroom Katherine E. Hagerman, 2010 Sustainable Design Through Process Integration Mahmoud M. El-Halwagi, 2011-10-03 Whether you are a process engineer an industrial decision maker or a researcher this book provides you with comprehensive and easy to follow coverage of the fundamental concepts and practical techniques on the use of process integration to maximize the efficiency and sustainability of industrial processes 2D Nanomaterials Ram K. Gupta, 2022-05-26 2D nanomaterials have emerged as promising candidates for use in energy devices owing to their superior electrochemical properties surface area nanodevice integration multifunctionality printability and mechanical flexibility 2D Nanomaterials Chemistry and Properties covers basic concepts chemistries and properties along with theoretical considerations in designing new 2D nanomaterials especially for energy applications. This book Discusses the effect of doping structural variation phase and exfoliation on structural and electrochemical properties of 2D nanomaterials Presents

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Chemical Engineering Education ,2004 **Cumulated Index Medicus** ,1989 **International Symposium on** the Nucleolus, Its Structure and Function, Held at Montevideo, Uruguay, December 5-10, 1965 W. S. Vincent, O. L. Electrochemical Energy Conversion and Storage Systems for Future Sustainability Aneeya Kumar Samantara, Satyajit Ratha, 2020-11-16 This new volume discusses new and well known electrochemical energy harvesting conversion and storage techniques It provides significant insight into the current progress being made in this field and suggests plausible solutions to the future energy crisis along with approaches to mitigate environmental degradation caused by energy generation production and storage Topics in Electrochemical Energy Conversion and Storage Systems for Future Sustainability Technological Advancements address photoelectrochemical catalysis by ZnO hydrogen oxidation reaction for fuel cell application and miniaturized energy storage devices in the form of micro supercapacitors. The volume looks at the underlying mechanisms and acquired first hand information on how to overcome some of the critical bottlenecks to achieve long term and reliable energy solutions The detailed synthesis processes that have been tried and tested over time through rigorous attempts of many researchers can help in selecting the most effective and economical ways to achieve maximum output and efficiency without going through time consuming and complex steps The theoretical analyses and computational results corroborate the experimental findings for better and reliable energy solutions **Transition States of Biochemical Processes** R. Gandour, 2013-06-29 The transItIOn state theory has been from the point of its inception the most influential principle in the development of our knowledge of reaction mechanisms in solution It is natural that as the field of biochemical dynamics has achieved new levels of refinement its students have increasingly adopted the concepts and methods of transition state theory Indeed every dynamical problem of biochemistry finds its most elegant and economical statement in the terms of this theory Enzyme catalytic power for example derives from the interaction of enzyme and substrate structures in the transition state so that an understanding of this power must grow from a knowledge of these structures and interactions Similarly transition state interactions and the way in which they change as protein structure is altered constitute the pivotal feature upon which molecular evolution must turn The complete coupled dynamical system of the organism incorporating the transport of matter and energy as well as local chemical processes will eventually have to yield to a description of its component transition state structures and their energetic response characteristics even if the form of the description goes beyond present day transition state theory Finally the importance of biochemical effectors in medicine and agriculture carries the subject into the world of practical affairs in the use of transition state information for the construction of ultra potent biological agents **Energy Research Abstracts**, 1977-11 Semiannual with semiannual

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