NAME_	Period Date
	Predicting Reaction types S, D, SD, DB, C 62 NR
equatio	the type of reaction (or NR for no reaction,) complete the word equation, and write a balanced chemical on for each of the following. Make sure to include phase symbols. → Wildel State of model of model of the following and hydrochloric acid →
DD	$N(s) + N(q) \rightarrow N(c)_{s}(q) + H_{2}(g)$ 2. Calcium hydroxide and nitric acid \Rightarrow
NR	Ca(0H)= (aq) + 2HNO ₅ (q) -> Ca(NO ₃) ₃ +2H ₂ O 3. Aluminum and Magnesium →
S	Al (s) + Mg(s) X> contract 2 materials goth tend to 4. Mercury(II) and oxygen > Form contract
DD.	$2 + 9 (t) + 0_2 (9) \longrightarrow 2 + 9 0 (s)$ 5. Zinc chloride and hydrogen sulfide \Rightarrow
000	$ZnCl_2(aq) + H_2S(q) \longrightarrow ZnS(s) + 2HCL(aq)$ 6. Silver chloride and sodium nitrate \rightarrow
Pyroli	7. Sodium chlorate is heated > Na. NO3 (ag) X> both products are agricult most
DD	2 NaClO 2(s) A > 2 NaCl/m)+3O2 (a) bc a prompitate, N2O, 8. Barium nitrate and sodium chromate →
DD	Ba(NO3) 2 (ag) + Nacr O4 (ag) -> Ba CrO4(s) +2NaNO3
DD	NaBr (q) + AgNO $_3$ (aq) \rightarrow AgBr (s) + Na No $_3$ (aq) (eq)
D	$Ca_3(PO_4)_2(s) + Al_2(sO_4)_3(qq) \longrightarrow 2AlPO_4(s) + 3CoSO_4(s)$ 11. Zinc carbonate is heated \Rightarrow
	$Z_0 C_0 C_3 (s) \xrightarrow{A} Z_0 C_3 + C_0 C_3 + C_0 C_3 $
	$2K(s) + F_2(g) \longrightarrow 2KF(s)_{map ?}$ 13. Hydrogen and oxygen \Rightarrow
	$2H_2(g) + O_2(g) \longrightarrow 2H_2O(g)$ 14. (lead(II) and potassium chlorate \rightarrow
	Pb(s) + KClO ₃ (aq.) X >
	6Na(s) +2H3 POH (a) -> 2NO3POH (aq) + 3H2(g)

Writing Equations Predicting Reactions Answer Key

LL Leslie

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