In the space before the formula write I for ionic compounds, IH for hydrates, and A for acids. Then, write the name of the compound.

1. HCI 16. Ba(BrO<sub>3</sub>), 2. KCl \_\_\_\_ 3. NH,OH 18. HClO. \_\_\_ 5. HNO, \_\_\_\_ \_\_\_\_20. LiH \_\_\_ 6. Cu<sub>3</sub>O \_\_\_\_\_\_ \_\_\_\_21. MgBr, 7. NaOH \_\_\_\_\_\_ \_\_\_22. H<sub>2</sub>CO<sub>3</sub> 8. HF 23. Ba\_As\_ 10. CuSO, •5H,O 25. NaClO 12. CsOH 27. Cu,P 13. SnCl,•2H,O 28. K,C,O, \_\_\_14. H,SO<sub>4</sub> \_\_\_\_\_\_\_ 29. H,C<sub>4</sub>H<sub>4</sub>O<sub>7</sub> 15. FeCO<sub>1</sub> 30. Ca(MnO<sub>4</sub>)<sub>2</sub>

II. Write the formula for the following compounds:

31. manganese (II) hydroxide 32. silver chloride ammonium sulfide 30.30 3.4. phosphoric acid 3.5 barium hydroxide octahydrate 3.6 cesium carbonate 37. chromium (III) phosphide 38. cobalt (III) nitrite sulfurous acid 3.9. 40. hydrobromic acid 41. lead (IV) chloride

vanadium (III) dichromate

cadmium phosphate

nickel (III) chloride

gold (III) bromide

42.

43.

44. 45.

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Katrin Zwirglmaier

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