

## Chemical Formula Writing Worksheet Solutions

Write chemical formulas for the compounds in each box. The names are found by finding the intersection between the cations and anions. Example: The first box is the intersection between the "zinc" cation and the "chloride" anion, so you should write "ZnCl<sub>2</sub>", as shown.

|                 | <i>zinc</i>                                                    | <i>iron (II)</i>                                               | <i>iron (III)</i>                                              | <i>gallium</i>                                                 | <i>silver</i>                                     | <i>lead (IV)</i>                                               |
|-----------------|----------------------------------------------------------------|----------------------------------------------------------------|----------------------------------------------------------------|----------------------------------------------------------------|---------------------------------------------------|----------------------------------------------------------------|
| <i>chloride</i> | <b>ZnCl<sub>2</sub></b>                                        | <b>FeCl<sub>2</sub></b>                                        | <b>FeCl<sub>3</sub></b>                                        | <b>GaCl<sub>3</sub></b>                                        | <b>AgCl</b>                                       | <b>PbCl<sub>4</sub></b>                                        |
| <i>acetate</i>  | <b>Zn(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub></b> | <b>Fe(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub></b> | <b>Fe(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>3</sub></b> | <b>Ga(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>3</sub></b> | <b>Ag C<sub>2</sub>H<sub>3</sub>O<sub>2</sub></b> | <b>Pb(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>4</sub></b> |
| <i>nitrate</i>  | <b>Zn(NO<sub>3</sub>)<sub>2</sub></b>                          | <b>Fe(NO<sub>3</sub>)<sub>2</sub></b>                          | <b>Fe(NO<sub>3</sub>)<sub>3</sub></b>                          | <b>Ga(NO<sub>3</sub>)<sub>3</sub></b>                          | <b>AgNO<sub>3</sub></b>                           | <b>Pb(NO<sub>3</sub>)<sub>4</sub></b>                          |
| <i>oxide</i>    | <b>ZnO</b>                                                     | <b>FeO</b>                                                     | <b>Fe<sub>2</sub>O<sub>3</sub></b>                             | <b>Ga<sub>2</sub>O<sub>3</sub></b>                             | <b>Ag<sub>2</sub>O</b>                            | <b>PbO<sub>2</sub></b>                                         |
| <i>nitride</i>  | <b>Zn<sub>3</sub>N<sub>2</sub></b>                             | <b>Fe<sub>3</sub>N<sub>2</sub></b>                             | <b>FeN</b>                                                     | <b>GaN</b>                                                     | <b>Ag<sub>3</sub>N</b>                            | <b>Pb<sub>3</sub>N<sub>4</sub></b>                             |
| <i>sulfate</i>  | <b>ZnSO<sub>4</sub></b>                                        | <b>FeSO<sub>4</sub></b>                                        | <b>Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></b>              | <b>Ga<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></b>              | <b>Ag<sub>2</sub>SO<sub>4</sub></b>               | <b>Pb(SO<sub>4</sub>)<sub>2</sub></b>                          |

Write the formulas for the following compounds:

- 1) copper (II) chloride **CuCl<sub>2</sub>**
- 2) lithium acetate **LiC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>**
- 3) vanadium (III) selenide **V<sub>2</sub>Se<sub>3</sub>**
- 4) manganese (IV) nitride **Mn<sub>3</sub>N<sub>4</sub>**
- 5) beryllium oxide **BeO**
- 6) sodium sulfate **Na<sub>2</sub>SO<sub>4</sub>**
- 7) aluminum arsenide **AlAs**
- 8) potassium permanganate **KMnO<sub>4</sub>**
- 9) chromium (VI) cyanide **Cr(CN)<sub>6</sub>**
- 10) tin (II) sulfite **SnSO<sub>3</sub>**
- 11) vanadium (V) fluoride **VF<sub>5</sub>**
- 12) ammonium nitrate **NH<sub>4</sub>NO<sub>3</sub>**

# Writing Chemical Formulas Worksheet Key Practice Problems Level

**Judith Potter, Peter Cann**



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