

CHEMISTRY 11 – WORD EQUATIONS AND TYPES OF CHEMICAL REACTIONS WORKSHEET

- 1) Translate the following word equations into chemical symbols and balance the resulting equations. Do not include the phases.

(a) Potassium + water $\rightarrow$ potassium hydroxide + hydrogen



(b) Nitrogen + hydrogen $\rightarrow$ ammonia



- 2) In each of the following write a balanced chemical equation, including the phases.

(a) Liquid phosphoric acid reacts with aqueous barium hydroxide to give water and a precipitate of barium phosphate



(b) Solid aluminium oxide and aqueous sulphuric acid produce water and aqueous aluminium sulphate



(c) Nitrogen trifluoride gas and hydrogen gas react to form nitrogen gas and gaseous hydrogen fluoride



(d) Powdered sodium carbonate and aqueous hydrobromic acid react to form carbon dioxide gas, aqueous sodium bromide and water



- 3) Classify the following reactions.

Reaction	Type
$2\text{C}_2\text{H}_2 + 5\text{O}_2 \rightarrow 4\text{CO}_2 + 2\text{H}_2\text{O}$	Combustion
$\text{Mg} + \text{CuSO}_4 \rightarrow \text{MgSO}_4 + \text{Cu}$	Single Replacement
$4\text{Na} + \text{O}_2 \rightarrow 2\text{Na}_2\text{O}$	Synthesis
$\text{Sn}(\text{OH})_4 + 4\text{HBr} \rightarrow 4\text{H}_2\text{O} + \text{SnBr}_4$	Neutralization
$2\text{Fe}(\text{NO}_3)_3 + 3\text{MgS} \rightarrow \text{Fe}_2\text{S}_3 + 3\text{Mg}(\text{NO}_3)_2$	Double Replacement
$2\text{N}_2\text{O} \rightarrow 2\text{N}_2 + \text{O}_2$	Decomposition

Writing Chemical Equations And Answers

Chao Zhang



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Table of Contents Writing Chemical Equations And Answers

1. Understanding the eBook Writing Chemical Equations And Answers
 - The Rise of Digital Reading Writing Chemical Equations And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Writing Chemical Equations And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Writing Chemical Equations And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Writing Chemical Equations And Answers
 - Personalized Recommendations
 - Writing Chemical Equations And Answers User Reviews and Ratings
 - Writing Chemical Equations And Answers and Bestseller Lists

5. Accessing Writing Chemical Equations And Answers Free and Paid eBooks
 - Writing Chemical Equations And Answers Public Domain eBooks
 - Writing Chemical Equations And Answers eBook Subscription Services
 - Writing Chemical Equations And Answers Budget-Friendly Options
6. Navigating Writing Chemical Equations And Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Writing Chemical Equations And Answers Compatibility with Devices
 - Writing Chemical Equations And Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Writing Chemical Equations And Answers
 - Highlighting and Note-Taking Writing Chemical Equations And Answers
 - Interactive Elements Writing Chemical Equations And Answers
8. Staying Engaged with Writing Chemical Equations And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Writing Chemical Equations And Answers
9. Balancing eBooks and Physical Books Writing Chemical Equations And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Writing Chemical Equations And Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Writing Chemical Equations And Answers
 - Setting Reading Goals Writing Chemical Equations And Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Writing Chemical Equations And Answers
 - Fact-Checking eBook Content of Writing Chemical Equations And Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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