Solutions Chemistry WebQuest

Directions:

- Go to http://www.wwnorton.com/college/chemistry/gilbert2/contents/ch04/studyplan.asp
- Under Chapter 4 ChemTours click the link that matches with the following headings.
- If it asks you to click on anything, follow the instructions. Click on any Concept Questions, Real-World Connections, or Science Connections to get more information.

Molarity

- Understanding the chemistry of liquid mixtures requires knowledge of amount or of the compounds in the solution.
- The concentration of solute is most usefully given in
- 3. What is the equation for this?
- 4. How can you determine the number of moles of something?
- Complete questions 1-7. Record your answers.
- Calculating the molar concentration, or _____, of a solution is probably the ______ calculation done by ______ students and researchers.

Dillutions

- Dilution occurs when additional is added to the solution.
- The molecules are chemically unchanged during dilution.
- When might you use a dilution? (read-world connections)
- 10. What is the equation used to solve dilution problems?
- 11. Both sides of the equation have exactly the same # of of solute.
- 12. For solutions of the same compound in the same type of container, the _____ of the color is an indicator of solute_____.

 13. The palest solution is the _____.
- Answer questions 1& 2
- Give 2 examples of when dilutions are used in science.

Migration of Ions

- 16. Some aqueous solutions will conduct _____ and some will not.
- The flow of electricity is increased greatly when ______ ions travel to terminals & aid in _____ transfer.
- water is a poor conductor, but or water may contain ions that may cause electricity.
- 19. Why should you not go swimming outside during a thunderstorm?
- In , ions can move freely and thereby conduct electricity.
- 21. What happens when a voltage is applied to Sodium Chloride Solution?
- 22. Why was the light bulb faint from the Acetic Acid solution?
- Strong acid & bases, and ionic compounds are
- & ____ are nonelectrolytes.
- 25. What types of ions would be found around the positive electrode?
- 26. What happens when 2 weak electrolytes are mixed?
- 27. Conductivity is another way to measure ______ in solution

Saturated Solutions

- 28. A saturated solution is one that contains the _____ concentration of _____.
 29. A _____ is created when a substance is _____ in a ____.
- When a solvent has reached its capacity and can dissolve no additional solute, it is said to be
- 31. Give an example of a saturated solution.
- 32. What happens to compounds in an unsaturated solution?
- 33. In a saturated solution, excess molecules do not dissolve, but form a on the bottom of the beaker.
- 34. What is happening to the molecule/ions in the beaker (section 3 of 4)?
- 35. What happens to solubility as temperature increases?

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G Thomas

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